

Alberta Exam Chem 30 unit 4 Version 2 ANSWER KEY

1. In the acid-base reaction



- A) reactants are favoured
 B) $\text{HC}_6\text{H}_5\text{O}_2^-(\text{aq})$ is acting as an acid
 C) $\text{SO}_4^{2-}(\text{aq})$ is a proton acceptor
 D) $\text{HSO}_4^-(\text{aq})$ and $\text{HC}_6\text{H}_5\text{O}_2^-(\text{aq})$ are a conjugate acid-base pair

Incorrect. Your answer=A, Correct answer=C

Explanation: In the above reaction, the conjugate acid base pair is HSO_4^- (acting as an acid, proton donor) and SO_4^{2-} (conjugate base, proton acceptor). Choice C is a correct answer.

- 2.

Tooth decay results from the dissolving of tooth enamel, or hydroxyapatite, $\text{Ca}_5(\text{PO}_4)_3\text{OH}(\text{s})$. Bacteria in the presence of sugars in the mouth change the pH of the mouth which causes the following equilibrium:



When sugar ferments on teeth and gums, the mouth becomes more acidic. Tooth enamel

- A) is formed as the concentration of the hydroxide ion decreases and the equilibrium shifts to the left
 B) is formed as the concentration of the hydroxide ion increases and the equilibrium shifts to the left
 C) breaks down as the concentration of the hydroxide ion decreases and the equilibrium shifts to the right
 D) breaks down as the concentration of the hydroxide ion increases and the equilibrium shifts to the right

Correct. Your answer=C, Correct answer=C

Explanation: Using your common sense and remembering your last visit to the dentist can help you out here. $\text{Ca}_5(\text{PO}_4)_3\text{OH}(\text{s})$, tooth enamel, breaks down in an acid environment. The H^+ ions would neutralize the OH^- ions and the decrease would shift the equilibrium to the right, breaking down the enamel. Only Choice C is correct.

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Shasha Hu



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1. Understanding the eBook Alberta Science 30 Answer Key
 - The Rise of Digital Reading Alberta Science 30 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Alberta Science 30 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Alberta Science 30 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Alberta Science 30 Answer Key
 - Personalized Recommendations
 - Alberta Science 30 Answer Key User Reviews and Ratings
 - Alberta Science 30 Answer Key and Bestseller Lists
5. Accessing Alberta Science 30 Answer Key Free and Paid eBooks
 - Alberta Science 30 Answer Key Public Domain eBooks
 - Alberta Science 30 Answer Key eBook Subscription Services
 - Alberta Science 30 Answer Key Budget-Friendly Options

6. Navigating Alberta Science 30 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Alberta Science 30 Answer Key Compatibility with Devices
 - Alberta Science 30 Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Alberta Science 30 Answer Key
 - Highlighting and Note-Taking Alberta Science 30 Answer Key
 - Interactive Elements Alberta Science 30 Answer Key
8. Staying Engaged with Alberta Science 30 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Alberta Science 30 Answer Key
9. Balancing eBooks and Physical Books Alberta Science 30 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Alberta Science 30 Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Alberta Science 30 Answer Key
 - Setting Reading Goals Alberta Science 30 Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Alberta Science 30 Answer Key
 - Fact-Checking eBook Content of Alberta Science 30 Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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