

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

a. Write the balanced chemical equation.

b. Identify what is given (with units) and what you want to find (with units).

c. Use coefficients from balanced equation to determine mole ratio.

d. Show set up (organize it!).

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?



Before $\underline{\quad 2 \quad}$ $\underline{\quad \text{ex} \quad}$ $\underline{\quad 0 \quad}$ $\underline{\quad 0 \quad}$ 3.0 mol O₂

Change $\underline{\quad -2 \quad}$ $\underline{\quad -3 \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

After $\underline{\quad 0 \quad}$ $\underline{\quad \text{ex} \quad}$ $\underline{\quad 2 \quad}$ $\underline{\quad 2 \quad}$

2. Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?



Before 0.5 $\underline{\quad \text{ex} \quad}$ 0 0 1.50mol HCL

Change -0.5 -1.5 +0.5 +0.5

After 0 $\underline{\quad \text{ex} \quad}$ 0.5 0.75

3. Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?



Before $\underline{\quad \text{ex} \quad}$ 1.25 0 2.50 mole

Change -3.75 -1.25 +2.5

After $\underline{\quad \text{ex} \quad}$ 0 2.5

Modeling Chemistry Unit 8 Mole Relationships Answers

Clemens Wendtner



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Applied Hydrogeochemistry Alan L. Mayo, 2025-12-23 Groundwater flow systems can be evaluated by solute and isotopic methods both powerful tools for students and professionals in hydrogeology This accessible textbook provides foundational knowledge and promotes a comprehensive understanding users need to fully evaluate both the solute and isotopic aspects of groundwater It focuses on groundwater chemistry and discusses the most important solutes stable and radiogenic isotopes and groundwater age dating topics often inadequately treated or ignored in other texts The reader will find chapters on equilibria gases kinetics organic compounds and the most important species carbon sulfur nitrogen silica noble gas and chloride and stable isotopes groundwater dating graphical methods and geothermometers Features Presents a balanced approach between solutes and isotopes the first book of its kind to do so Explains in detail how to apply stable isotopic data to help solve groundwater problems Describes the appropriate groundwater dating tools for both young and old groundwater and how to apply the tools Provides the reader with the background knowledge necessary to conceptualize the geochemical evolution of a groundwater system before using a geochemical modeling program Examines how geothermometers and graphical methods can be used to illustrate chemical evolution This textbook is an excellent resource for upper level undergraduates and graduate students taking courses in groundwater chemistry that are taught in departments of geology physical geography environmental science and environmental engineering as well as for researchers and working professionals involved with water and groundwater

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