

## Chemistry – Unit 10 Worksheet 3

1. Element X has two natural isotopes. The isotope with a mass number of 6 has a relative abundance of 7.5%. The isotope with a mass number of 7 has a relative abundance of 92.5%. Determine the average molar mass for the element from these figures. What is the true identity and atomic number of element X?
2. The element copper is found to contain the naturally occurring isotopes  ${}^{63}\text{Cu}$  and  ${}^{65}\text{Cu}$ . The relative abundances are 69.1% and 30.9% respectively. Calculate the average molar mass of copper.
3. Uranium has three isotopes with the following percent abundances:  ${}^{234}\text{U}$  (0.0058%),  ${}^{235}\text{U}$  (0.71%),  ${}^{238}\text{U}$  (99.23%). What do you expect the molar mass of uranium to be in whole numbers? Why?
4. A sample of silver as it occurs in nature is 52.0% of isotope  ${}^{107}\text{Ag}$  and 48.0% of isotope  ${}^{109}\text{Ag}$ . What is the average molar mass of silver? (Compare your result with the value given in the periodic table).

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**Robert R. Redfield**



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## **Table of Contents Modeling Chemistry U9 Ws 2 V2 Answers**

1. Understanding the eBook Modeling Chemistry U9 Ws 2 V2 Answers
  - The Rise of Digital Reading Modeling Chemistry U9 Ws 2 V2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry U9 Ws 2 V2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry U9 Ws 2 V2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry U9 Ws 2 V2 Answers
  - Personalized Recommendations
  - Modeling Chemistry U9 Ws 2 V2 Answers User Reviews and Ratings
  - Modeling Chemistry U9 Ws 2 V2 Answers and Bestseller Lists
5. Accessing Modeling Chemistry U9 Ws 2 V2 Answers Free and Paid eBooks
  - Modeling Chemistry U9 Ws 2 V2 Answers Public Domain eBooks
  - Modeling Chemistry U9 Ws 2 V2 Answers eBook Subscription Services
  - Modeling Chemistry U9 Ws 2 V2 Answers Budget-Friendly Options
6. Navigating Modeling Chemistry U9 Ws 2 V2 Answers eBook Formats

- ePub, PDF, MOBI, and More
  - Modeling Chemistry U9 Ws 2 V2 Answers Compatibility with Devices
  - Modeling Chemistry U9 Ws 2 V2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modeling Chemistry U9 Ws 2 V2 Answers
    - Highlighting and Note-Taking Modeling Chemistry U9 Ws 2 V2 Answers
    - Interactive Elements Modeling Chemistry U9 Ws 2 V2 Answers
  8. Staying Engaged with Modeling Chemistry U9 Ws 2 V2 Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modeling Chemistry U9 Ws 2 V2 Answers
  9. Balancing eBooks and Physical Books Modeling Chemistry U9 Ws 2 V2 Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modeling Chemistry U9 Ws 2 V2 Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modeling Chemistry U9 Ws 2 V2 Answers
    - Setting Reading Goals Modeling Chemistry U9 Ws 2 V2 Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modeling Chemistry U9 Ws 2 V2 Answers
    - Fact-Checking eBook Content of Modeling Chemistry U9 Ws 2 V2 Answers
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements
    - Interactive and Gamified eBooks

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