

9

Thermodynamics

1. Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
- (a) ΔS is negative and therefore, ΔH should be highly positive
(b) ΔS is negative and therefore, ΔH should be highly negative
(c) ΔS is positive and therefore, ΔH should be negative
(d) ΔS is positive and therefore, ΔH should also be highly positive
2. For the reaction, $X_2O_4(l) \longrightarrow 2XO_2(g)$, $\Delta U = 2.1 \text{ kcal}$, $\Delta S = 20 \text{ cal K}^{-1}$ at 300 K. Hence, ΔG is [2014]
- (a) 2.7 kcal (b) - 2.7 kcal
(c) 9.3 kcal (d) - 9.3 kcal
3. For a given exothermic reaction, K_p and K_p' are the equilibrium constants at temperatures T_1 and T_2 , respectively. Assuming that heat of reaction is constant in temperature range between T_1 and T_2 , it is readily observed that [2014]
- (a) $K_p > K_p'$
(b) $K_p < K_p'$
(c) $K_p = K_p'$
(d) $K_p = \frac{1}{K_p'}$
4. A reaction having equal energies of activation for forward and reverse reactions has [2013]
- (a) $\Delta S = 0$ (b) $\Delta G = 0$
(c) $\Delta H = 0$ (d) $\Delta H = \Delta G = \Delta S = 0$
5. In which of the following reactions, standard reaction entropy changes (ΔS°) is positive and standard Gibbs energy change (ΔG°) decreases sharply with increasing temperature? [2012]
- (a) $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$
(b) $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
(c) $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
(d) $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
6. The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
- (a) 10.52 cal/mol K (b) 21.04 cal/mol K
(c) 5.260 cal/mol K (d) 0.526 cal/mol K
7. Standard enthalpy of vaporisation $\Delta_{\text{vap}}H^\circ$ for water at 100°C is 40.66 kJ mol⁻¹. The internal energy of vaporisation of water at 100°C (in kJ mol⁻¹) is (assume water vapour to behave like an ideal gas). [2012]
- (a) + 37.56 (b) - 43.76
(c) + 43.76 (d) + 40.66
8. If the enthalpy change for the transition of liquid water to steam is 30 kJ mol⁻¹ at 27°C, the entropy change for the process would be [2011]
- (a) 1.0 J mol⁻¹K⁻¹
(b) 0.1 J mol⁻¹K⁻¹
(c) 100 J mol⁻¹K⁻¹
(d) 10 J mol⁻¹K⁻¹

Conceptual Physics Thermodynamics Review Answers

SA Adler



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Table of Contents Conceptual Physics Thermodynamics Review Answers

1. Understanding the eBook Conceptual Physics Thermodynamics Review Answers
 - The Rise of Digital Reading Conceptual Physics Thermodynamics Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Conceptual Physics Thermodynamics Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Conceptual Physics Thermodynamics Review Answers
 - Personalized Recommendations
 - Conceptual Physics Thermodynamics Review Answers User Reviews and Ratings
 - Conceptual Physics Thermodynamics Review Answers and Bestseller Lists
5. Accessing Conceptual Physics Thermodynamics Review Answers Free and Paid eBooks
 - Conceptual Physics Thermodynamics Review Answers Public Domain eBooks
 - Conceptual Physics Thermodynamics Review Answers eBook Subscription Services
 - Conceptual Physics Thermodynamics Review Answers Budget-Friendly Options

6. Navigating Conceptual Physics Thermodynamics Review Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Conceptual Physics Thermodynamics Review Answers Compatibility with Devices
 - Conceptual Physics Thermodynamics Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Conceptual Physics Thermodynamics Review Answers
 - Highlighting and Note-Taking Conceptual Physics Thermodynamics Review Answers
 - Interactive Elements Conceptual Physics Thermodynamics Review Answers
8. Staying Engaged with Conceptual Physics Thermodynamics Review Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Conceptual Physics Thermodynamics Review Answers
9. Balancing eBooks and Physical Books Conceptual Physics Thermodynamics Review Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Conceptual Physics Thermodynamics Review Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Conceptual Physics Thermodynamics Review Answers
 - Setting Reading Goals Conceptual Physics Thermodynamics Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Conceptual Physics Thermodynamics Review Answers
 - Fact-Checking eBook Content of Conceptual Physics Thermodynamics Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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