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REPORT SHEET | LAB  
**Compounds and  
 Their Bonds** | **9**

**A. Ions: Transfer of Electrons**

Element	Atomic Number	1. Electron Configuration of Atom	2. Electron-Dot Symbol	3. Loss or Gain of Electrons	4. Electron Configuration of Ion	5. Ionic Charge	6. Symbol of Ion	7. Name of Ion
Sodium	11	$1s^2 2s^2 2p^6 3s^1$	Na <sup>•</sup>	lose 1 e <sup>-</sup>	$1s^2 2s^2 2p^6$	1+	Na <sup>+</sup>	Sodium
Nitrogen	7	$1s^2 2s^2 2p^3$	$:\ddot{N}:$	gains 3e <sup>-</sup>	$1s^2 2s^2 2p^6$	3-	N <sup>3-</sup>	nitride
Aluminum	13	$1s^2 2s^2 2p^6 3s^2 3p^1$	$\cdot\text{Al}\cdot$	lose 3e <sup>-</sup>	$1s^2 2s^2 2p^6$	3+	Al <sup>3+</sup>	aluminum
Chlorine	17	$1s^2 2s^2 2p^6 3s^2 3p^5$	$:\ddot{Cl}\cdot$	gains 1e <sup>-</sup>	$1s^2 2s^2 2p^6 3s^2 3p^6$	1-	Cl <sup>-</sup>	chloride
Calcium	20	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	Ca <sup>••</sup>	lose 2e <sup>-</sup>	$1s^2 2s^2 2p^6 3s^2 3p^6$	2+	Ca <sup>2+</sup>	calcium
Oxygen	8	$1s^2 2s^2 2p^4$	$:\ddot{O}:$	gains 2e <sup>-</sup>	$1s^2 2s^2 2p^6$	2-	O <sup>2-</sup>	oxide

\*  $1s^2 2s^2 2p^6 3s^2 3p^6$

**B. Ionic Compounds and Formulas**

**1. Physical properties**

Compound	Appearance	Density	Melting Point
Sodium chloride, NaCl	white crystalline solid	2.17 g/cm <sup>3</sup>	1,474 °F

**2. Formulas of ionic compounds**

Name	Positive Ion	Negative Ion	Formula
Lithium iodide	Li <sup>+</sup>	I <sup>-</sup>	LiI
Aluminum oxide	Al <sup>3+</sup>	O <sup>2-</sup>	Al <sub>2</sub> O <sub>3</sub>
Calcium sulfide	Ca <sup>2+</sup>	S <sup>2-</sup>	CaS
Magnesium bromide	Mg <sup>2+</sup>	Br <sup>-</sup>	MgBr <sub>2</sub>
Potassium nitride	K <sup>+</sup>	N <sup>3-</sup>	K <sub>3</sub> N
Sodium fluoride	Na <sup>+</sup>	F <sup>-</sup>	NaF

# Compounds And Their Formulas Report Sheet Answers

**SA Dillow**



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