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web the carbon 14 nucleus is unstable it s radioactive it s going to undergo spontaneous decay it s going to try to get a better neutron to proton ratio so let s look at the nuclear equation which represents the spontaneous decay of carbon 14 so here is

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10 o 2 co 2 h 2 o nh 3 cuo cu n 2 h 2 o q10 5 5 balance the following equations fe s cl 2 g fecl 3 g c 4 h 10 o o 2 co 2 h 2 o as naoh na 3 aso 3 h 2

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web in alpha decay an alpha particle is ejected from an unstable nucleus so here s our unstable nucleus uranium 238 an alpha particle has the same composition as a helium nucleus we saw the helium nucleus in the previous video there are two protons in the helium nucleus and two neutrons

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web sep 12 2022 the half life of the nuclei and thus the material is t1 2 693 t 1 2 693 million years 35 a the decay constant is λ 1 99 10 5 s 1 λ 1 99 10 5 s 1 b since strontium 91 has an atomic mass of 90 90 g

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web complete the nuclear equation to find the atomic number of californium 24296 cm α 24 235 cf 01 n by firing heavier particles such as carbon or boron ions at the target materials heavier elements can be synthesised

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web nuclear equations represent the reactants and products in radioactive decay nuclear fission or nuclear fusion instead of chemical equations where it shows the different number of elements is conserved in a reaction in a nuclear reaction the atomic mass and proton number are conserved