



**LOW CARB
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Low Carb Recipes Compare Tutorial

Guillaume Favre



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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17.3.2.1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $ph = -\log 0.100 = 1.000$ ph $\log 0.100 = 1.000$ b titrant volume 12.50 ml

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web ab acid base titration curves purpose the purpose of this lab is to determine the shape of acid base titration curves theoretically and experimentally and to use an acid base titration to measure the effectiveness of antacid tablets summary a titration curve is the plot of the pH value of a solution vs the amount of titrant added

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Volume of NaOH added (mL)	pH of the solution
4.6	4.6
4.6	5.9
5.9	5.9

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web background plot of the ph of a solution against the volume of titrant added is called a titration curve the ph can be measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base NaOH

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