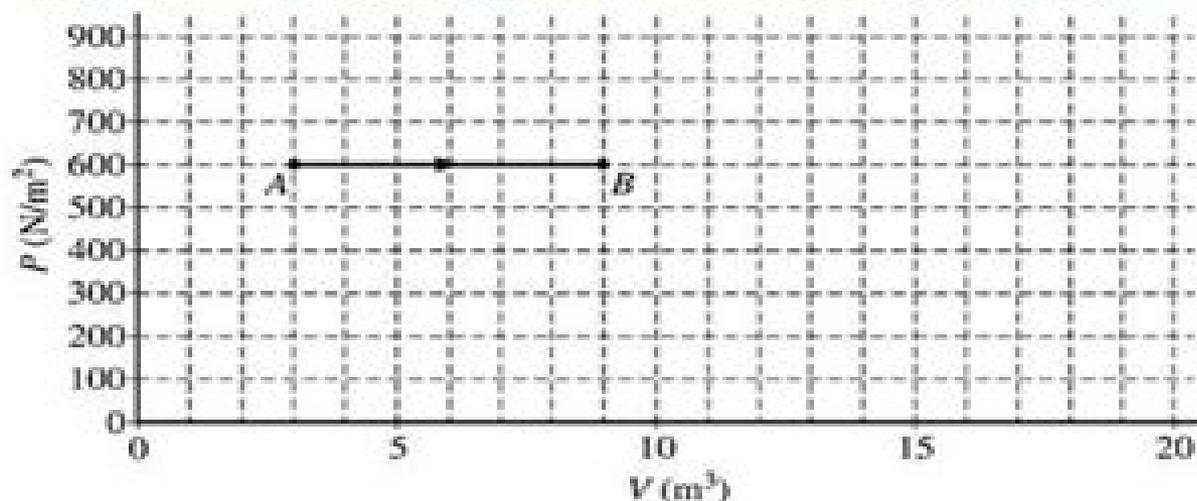


## 2004 AP<sup>®</sup> PHYSICS B FREE-RESPONSE QUESTIONS



5. (10 points)

The diagram above of pressure  $P$  versus volume  $V$  shows the expansion of 2.0 moles of a monatomic ideal gas from state  $A$  to state  $B$ . As shown in the diagram,  $P_A = P_B = 600 \text{ N/m}^2$ ,  $V_A = 3.0 \text{ m}^3$ , and  $V_B = 9.0 \text{ m}^3$ .

(a)

- i. Calculate the work done by the gas as it expands.
- ii. Calculate the change in internal energy of the gas as it expands.
- iii. Calculate the heat added to or removed from the gas during this expansion.

(b) The pressure is then reduced to  $200 \text{ N/m}^2$  without changing the volume as the gas is taken from state  $B$  to state  $C$ . Label state  $C$  on the diagram and draw a line or curve to represent the process from state  $B$  to state  $C$ .

(c) The gas is then compressed isothermally back to state  $A$ .

- i. Draw a line or curve on the diagram to represent this process.
- ii. Is heat added to or removed from the gas during this isothermal compression?

\_\_\_\_\_ added to                      \_\_\_\_\_ removed from

Justify your answer.

# Ap Physics 2010 Response Answers

**Greg Jacobs**

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