

# Covalent Bonding

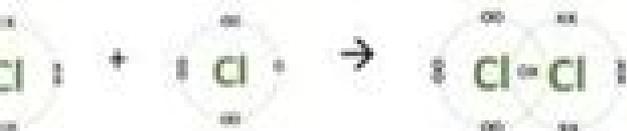
atoms need to get more electrons to have a full outer shell. They can do this in 2 ways:

**ionic bonding** - atoms lose or gain electrons, forming positive and negative ions which are then electrostatically attracted to each other (an ionic bond). This occurs between a metal and a non-metal.

**Covalent bonding** - atoms share a pair of electrons (a covalent bond). This occurs between non-metals only.

Bonding can only happen between non-metals - it does not involve metals.

atoms 'share' a pair of electrons. They do this by overlapping their outer shells so that the shared electrons are present in the shells of both atoms. This shared pair of electrons is a covalent bond.



note that as both atoms now possess the shared pair, both atoms have full outer shells of 8 electrons in their outer shell (in the covalent bond, 2 each on their outer shells).

hydrogen atoms are involved in covalent bonding, they will only have 2 electrons in their outer shell as the first shell can only contain 2 electrons. (remember their outer shell is still full).



## How covalent dot cross diagrams:

1. How many electrons do the atoms need to gain to have a full outer shell? This is how many bonds (dots) go in the centre.

2. How many electrons does the atom already have? If it has 7 electrons in its outer shell, it needs to gain 1 electron so it will form 1 covalent bond. This is how many dots (bond pairs) go in the centre.

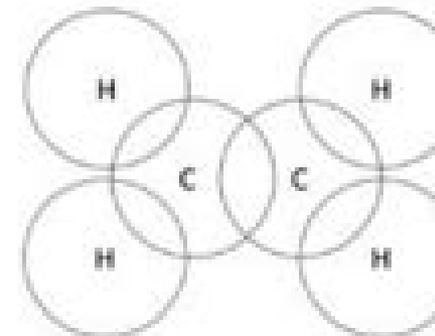
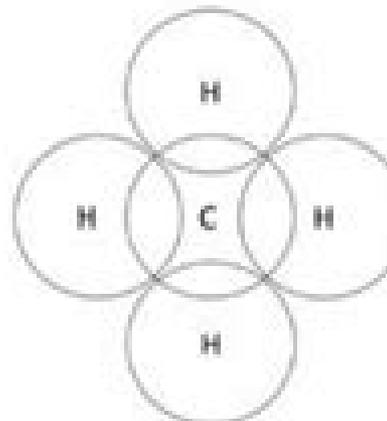
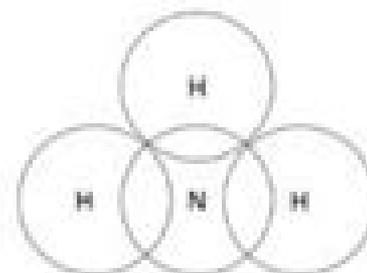
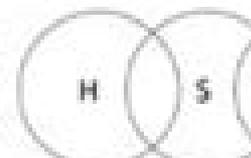
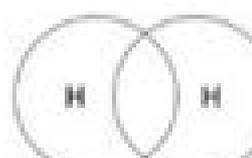
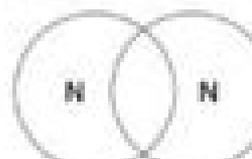
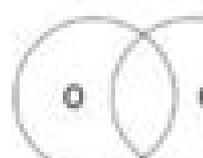
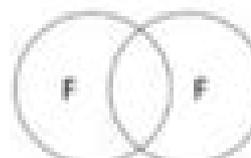
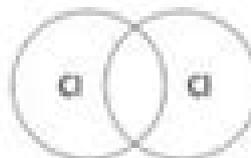


3. How many electrons do the atoms already have? If it has 4 electrons in its outer shell, it needs to gain 4 electrons so it will form 4 covalent bonds. Remember that each atom already 'uses' the 2 shared electrons in the centre.

4. How many electrons does the atom already have? If it has 2 electrons (from the shared pair), it then each need 6 more. Their outer shells are now full.



# Dot Cross Diagram Practice



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## **Table of Contents Dot Point Chemistry Answers Bonding**

1. Understanding the eBook Dot Point Chemistry Answers Bonding
  - The Rise of Digital Reading Dot Point Chemistry Answers Bonding
  - Advantages of eBooks Over Traditional Books
2. Identifying Dot Point Chemistry Answers Bonding
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Dot Point Chemistry Answers Bonding
  - User-Friendly Interface
4. Exploring eBook Recommendations from Dot Point Chemistry Answers Bonding

- Personalized Recommendations
  - Dot Point Chemistry Answers Bonding User Reviews and Ratings
  - Dot Point Chemistry Answers Bonding and Bestseller Lists
5. Accessing Dot Point Chemistry Answers Bonding Free and Paid eBooks
    - Dot Point Chemistry Answers Bonding Public Domain eBooks
    - Dot Point Chemistry Answers Bonding eBook Subscription Services
    - Dot Point Chemistry Answers Bonding Budget-Friendly Options
  6. Navigating Dot Point Chemistry Answers Bonding eBook Formats
    - ePub, PDF, MOBI, and More
    - Dot Point Chemistry Answers Bonding Compatibility with Devices
    - Dot Point Chemistry Answers Bonding Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Dot Point Chemistry Answers Bonding
    - Highlighting and Note-Taking Dot Point Chemistry Answers Bonding
    - Interactive Elements Dot Point Chemistry Answers Bonding
  8. Staying Engaged with Dot Point Chemistry Answers Bonding
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Dot Point Chemistry Answers Bonding
  9. Balancing eBooks and Physical Books Dot Point Chemistry Answers Bonding
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Dot Point Chemistry Answers Bonding
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Dot Point Chemistry Answers Bonding
    - Setting Reading Goals Dot Point Chemistry Answers Bonding
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Dot Point Chemistry Answers Bonding

- Fact-Checking eBook Content of Dot Point Chemistry Answers Bonding
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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