

KEY

1. Complete and classify the following partial equations

- a) $1 \text{ Fe} + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Fe SO}_4 + 3 \text{ H}_2$ (single repl)
- b) $\text{NaI} + \text{Cl}_2 \rightarrow$
- c) $1 \text{ Zn} + 1 \text{ Pb(NO}_3)_2 \rightarrow 1 \text{ Zn(NO}_3)_2 + 1 \text{ Pb}$ (single repl)
- d) $\text{Al(OH)}_3 + \text{HCl} \rightarrow$
- e) $2 \text{ Cu} + \text{O}_2 \rightarrow 2 \text{ CuO}$ $\text{4 Cu} + 1 \text{ O}_2 \rightarrow 2 \text{ Cu}_2\text{O}$ (synthesis)
- f) $\text{Al}_2\text{O}_3 \rightarrow$
- g) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ (synth)
- h) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- i) $1 \text{ Mg} + 2 \text{ HNO}_3 \rightarrow 1 \text{ Mg(NO}_3)_2 + 1 \text{ H}_2$ (s.r.)
- j) $\text{Na} + \text{H}_2\text{O} \rightarrow$
- k) $3 \text{ NaOH} + \text{H}_3\text{PO}_4 \rightarrow 1 \text{ Na}_3\text{PO}_4 + 3 \text{ H}_2\text{O}$ (acid base)
- l) $\text{H}_2\text{SO}_4 + \text{Al(OH)}_3 \rightarrow$
- m) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ (combustion)
- n) $\text{AlCl}_3 \rightarrow$
- o) $2 \text{ KOH} + \text{H}_2\text{SO}_4 \rightarrow 1 \text{ K}_2\text{SO}_4 + 2 \text{ H}_2\text{O}$ (acid base)
- p) $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow$
- q) $1 \text{ NaOH} + \text{HCl} \rightarrow 1 \text{ NaCl} + 1 \text{ H}_2\text{O}$ (acid base)
- r) $\text{NaHCO}_3 + \text{H}_2\text{SO}_4 \rightarrow$
- s) $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ (combustion)
- t) $\text{H}_2\text{O} \rightarrow$
- u) $2 \text{ HNO}_3 + \text{CuCO}_3 \rightarrow 1 \text{ Cu(NO}_3)_2 + 1 \text{ CO}_2$ (s.r.)
- v) $\text{CH}_4 + \text{O}_2 \rightarrow$
- w) $1 \text{ Pb} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ PbSO}_4 + 1 \text{ H}_2$ (s.r.)
- x) $\text{KI} + \text{Br}_2 \rightarrow$
- y) $1 \text{ Ba(OH)}_2 + 2 \text{ HNO}_3 \rightarrow 1 \text{ Ba(NO}_3)_2 + 2 \text{ H}_2\text{O}$ (acid base)
- z) $\text{Pb} + \text{Cu(NO}_3)_2 \rightarrow$



3 each

2 each

3

4

1/2/1/1/1

1/2/1

1/2/1/1

5

2. Given: $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$ calculate the mass of oxygen gas produced if 143 g of NaClO_3 is decomposed

65.4g

3. Calculate the mass of oxygen gas required to burn 242 g of wax ($\text{C}_{21}\text{H}_{44}$)

254g

4. If 1.40 moles of butane (C_4H_{10}) is burned

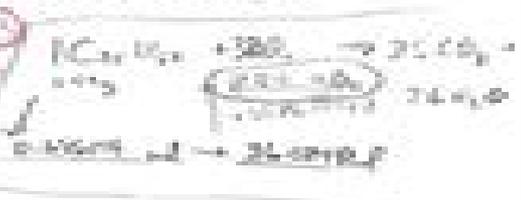
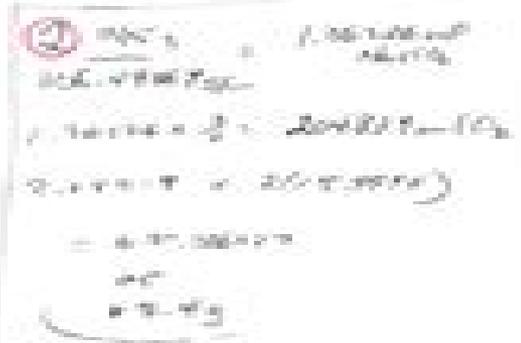
- a) calculate the number of moles of oxygen gas required (9.8 moles)
- b) calculate the number of moles of carbon dioxide produced (5.6 moles)
- c) calculate the mass of water produced (10.8 moles H_2O)

5. Metallic aluminum reacts with zinc nitrate to produce metallic zinc and aluminum nitrate

- a) Calculate the mass of aluminum that will react with 1.30 moles of zinc nitrate. (27.4g Al)
- b) Calculate the mass of zinc that would be produced from 50.0 g of aluminum and sufficient zinc nitrate. (192 g Zn)

6. Given: $\text{Cu} + 2 \text{ AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{ Ag}$

- a) How many grams of Cu are required to react with 3.00 g of silver nitrate? (0.544g Cu)
- b) How many grams of silver are produced from 3.00 g of silver nitrate and sufficient Cu? (1.70g Ag)



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Huangqi Zhang



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web 10 15 06 dr vince scialli bsc 1086 exam 2 test bank cardiovascular system 1 the portion of the intrinsic cardiac conduction system located in the interatrial septum is the a av node b sa node c tricuspid node d purkinje fibers e none of the above 2 which is are true about extrinsic stimulation of the heart a

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web nov 9 2023 the circulatory system or cardiovascular system is a system in which blood circulates throughout the body this organ includes blood blood vessels and the heart the main function of this system is to provide oxygen hormones to muscles and nutrients throughout the body