



## Contact

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### Address

New York, USA

### Phone

+1 2345 6789

### Email

max.johnson@email.com

## Skills

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- Figma
- Adobe XD
- Sketch
- InVision
- Photoshop

## Languages

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- English
- French
- Russian

## Hobbies

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- Photography
- Football
- Cooking

# MAX JOHNSON

## UX Designer

## Profile

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Experienced UX Designer specializing in user research, interaction design, and prototyping. Committed to crafting intuitive and visually compelling digital experiences that captivate users. A problem-solver who combines creativity with data-driven insights to drive innovation.

## Work Experience

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### Senior UX Designer

ABC Tech Solutions – New York

Jan 2020 – Dec 2022

- Led UX team for mobile banking app, boosting user engagement by 30% in six months.
- Conducted in-depth user research, reducing e-commerce cart abandonment by 25% through redesign.

### UX Designer

XYZ Design Agency – Texas

Jan 2018 – Dec 2019

- Collaborated with diverse clients to align designs with business objectives, achieving a 15% boost in healthcare app user satisfaction.
- Conducted impactful user tests on a financial site, pinpointing pain points and driving a 20% surge in conversion rates.

## Education

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### Master of Science in HCI

New York University – New York

Jan 2016 – Dec 2018

### Bachelor of Science in UX Design

University of Washington - Washington

Jan 2011 – Dec 2015

# Resume Template Top Open Now

**Michael Brown**



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$   $\Delta H = 94.5 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$   $\Delta H = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$   $\Delta H = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$   $\Delta H = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$   $\Delta H = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta H = 285 \text{ kJ/mole}$

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web mar 13 2023 be able to use hess s law to calculate reaction enthalpies first law of thermodynamics energy is the ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions  $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$   $\text{H}_2\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq})$   $\text{H}_2\text{O}(\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}))$

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