

## Chapter 8

# Organic Chemistry – Some Basic Principles and Techniques

### Question 8.1:

What are hybridisation states of each carbon atom in the following compounds?

$\text{CH}_2 = \text{C} = \text{O}$ ,  $\text{CH}_3\text{CH} = \text{CH}_2$ ,  $(\text{CH}_3)_2\text{CO}$ ,  $\text{CH}_2 = \text{CHCN}$ ,  $\text{C}_6\text{H}_6$

#### Answer 8.1:

- (i)  $\overset{1}{\text{C}}\text{H}_2 = \overset{2}{\text{C}} = \overset{3}{\text{O}}$   
C-1 is  $sp^2$  hybridised.  
C-2 is  $sp$  hybridised.
- (ii)  $\overset{1}{\text{C}}\text{H}_3 - \overset{2}{\text{C}}\text{H} = \overset{3}{\text{C}}\text{H}_2$   
C-1 is  $sp^3$  hybridised.  
C-2 is  $sp^2$  hybridised.  
C-3 is  $sp^2$  hybridised.
- (iii)  $\overset{1}{\text{C}}\text{H}_3 - \overset{2}{\text{C}} = \overset{3}{\text{O}} - \overset{4}{\text{C}}\text{H}_3$   
C-1 is  $sp^3$  hybridised.
- (iv)  $\overset{1}{\text{C}}\text{H}_2 = \overset{2}{\text{C}}\text{H} - \overset{3}{\text{C}} = \text{N}$   
C-2 is  $sp^2$  hybridised.  
C-3 is  $sp$  hybridised.
- (v)  $\text{C}_6\text{H}_6$   
All the 6 carbon atoms in benzene are  $sp^2$  hybridised.

### Question 8.2:

Indicate the  $\sigma$  and  $\pi$  bonds in the following molecules :

$\text{C}_6\text{H}_6$

$\text{C}_6\text{H}_{12}$

$\text{CH}_2\text{Cl}_2$

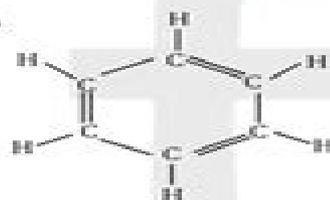
$\text{CH}_2 = \text{C} = \text{CH}_2$

$\text{CH}_3\text{NO}_2$

$\text{HCONHCH}_3$

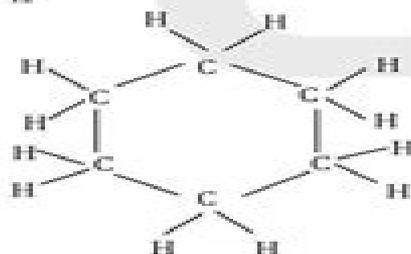
#### Answer 8.2:

(i)  $\text{C}_6\text{H}_6$



There are six C-C sigma ( $\sigma_{C-C}$ ) bonds, six C-H sigma ( $\sigma_{C-H}$ ) bonds and three C=C pi ( $\pi_{C=C}$ ) resonating bonds in this compound.

(ii)  $\text{C}_6\text{H}_{12}$



There are six C-C sigma ( $\sigma_{C-C}$ ) bonds and twelve C-H sigma ( $\sigma_{C-H}$ ) bonds in this compound.

# Chemistry Chapter 8

**VM Jensen**



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