

Atomic structure

Label the following diagram with the **correct** proton, neutron, electron and nucleus.



Carbon atom

2. Now complete the table below, identifying the charges, mass and place found for each of the sub-atomic particles.

Sub-atomic particle	Place found	Mass	Charge
proton			
neutron			
electron			

3. Why are atoms neutral overall?

.....

.....

Atomic structure 2

Use the information in the periodic table to help you complete the table below:

Element	Protons	Neutrons	Electrons
Carbon			
Sodium			
Magnesium			
Isotope			
Bromine			
Hydrogen			
Aluminium			
Argon			

3. What is an isotope?

.....

.....

Calculating Relative Atomic Mass "A_r"

Calculate the relative atomic mass of the following elements, given their mass and abundance.

a) Bromine, consisting of 50% Br-79 and 50% Br-81

.....

b) Copper, consisting of 69.09% Cu-63 and 30.91% Cu-65

.....

c) Silver, consisting of 97.54% Ag-107 and 2.46% Ag-109

.....

The Periodic Table

1. What element is found in:

- a) Group 2 Period 4 _____
- b) Group 1 Period 1 _____
- c) Group 7 Period 2 _____
- d) Group 8 Period 3 _____

2. What group and period are the following elements found in?

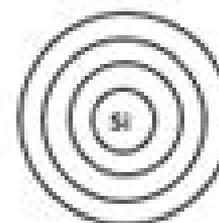
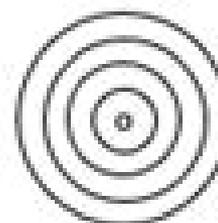
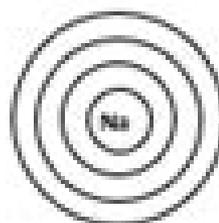
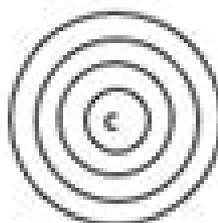
- a) Lithium _____
- b) Phosphorus _____
- c) Iron _____
- d) Rubidium _____

Electron structure

1. How many electrons can each shell hold?

- a) 1st shell _____
- b) 2nd shell _____
- c) 3rd shell _____
- d) 4th shell _____

2. Complete the electron structure diagrams for the following elements by drawing electrons on the shells. Under each diagram, WRITE the electron configuration, e.g. 2,8,1



The Periodic Table

1. Shade the metals in blue
2. Shade the non-metals in red
3. Shade the metalloids in green
4. What is special about the properties of metalloids? _____

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Albert A Gayle



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