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Title: What Is Chemical Reaction?

Introduction: The goal of this experiment is to determine chemical properties of hydrochloric acid and copper chloride; identify the evidence that we can gather to find out which chemical change that occurred, whether in nature or laboratory; if the law of conservation of mass applies to a sample chemical reaction in one of these series. A chemical change is defined as a variable composition and properties of substance. There are many types of observable changes such as formation of solid precipitate upon mixing two solutions; release of gas bubbles that are not due to a physical change; a color change that does not result from dilution or color mixing; or a temperature change that is not caused by the external heating or cooling. For instance, a nail can become tarnished when it gets old as new substances begin to take the place of the old substances of the nail. Still, it is physically a nail but the "formula" differentiates. The nail now has been oxidized. If you are a cook, you are now a food scientist. Chemical changes don't have to occur in the lab all the time, they can actually be right there in your kitchen. Think of a simple breakfast recipe like scrambled eggs. Cooking an egg uses a chemical reaction because the heat polymerizes proteins in egg white, while the hydrogen and sulfur in the yolk can react to form hydrogen sulfide gas. Same thing happens with cooked meats, and cake.

Objective: to examine the chemical properties of hydrochloric acid and copper chloride, to identify the types of reactions that they undergo, and to determine the law of conservation of mass applies to a sample chemical reaction.

Hypothesis: as acid is a hazardous chemical, the samples might experience corroding or discolor as soaking in the hydrochloric solution. The samples may not react as dramatically to copper chloride as hydrochloric because it does not have an acid name.

Result:

Data Table A: Reactions of Hydrochloric Acid

Initial appearance of HCl Solutions: transparent, and clear liquid

Initial temperature of HCl solutions: 74° F

Reaction Well	Reagents	Observations
1	HCl + NaOH	After 1 minute: 72° F Litmus paper turns pink.
2	HCl + NaHCO ₃	Bubble frizzing, temperature decreases
3	HCl + AgNO ₃	Foam occurs in the beginning, then the solution

Chemical Reaction Lab Report Grade 10

Dr. Dennis Englin



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Chemistry (Teacher Guide) Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books. The teacher is one who guides students through the subject matter, helps each student stay on schedule and be organized, and is their source of accountability along the way. With that in mind, this guide provides additional help through the laboratory exercises as well as lessons, quizzes, and examinations that are provided along with the answers. The lessons in this study emphasize working through procedures and problem solving by learning patterns. The vocabulary is kept at the essential level. Practice exercises are given with their answers so that the patterns can be used in problem solving. These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college. Guided labs are provided to enhance instruction of weekly lessons. There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions. It is important to see the hand of God and His principles and wisdom as it plays out in chemistry. This course integrates what God has told us in the context of this study. Features: Each suggested weekly schedule has five easy-to-manage lessons that combine reading and worksheets. Worksheets, quizzes, and tests are perforated and three-hole punched; materials are easy to tear out, hand out, grade, and store. Adjust the schedule and materials needed to best work within your educational program. Space is given for assignments and dates. There is flexibility in scheduling. Adapt the days to your school schedule. Workflow: Students will read the pages in their book and then complete each section of the teacher guide. They should be encouraged to complete as many of the activities and projects as possible as well. Tests are given at regular intervals with space to record each grade. About the Author: DR DENNIS ENGLIN earned his bachelor's from Westmont College, his master of science from California State University, and his EdD from the University of Southern California. He enjoys teaching animal biology, vertebrate biology, wildlife biology, organismic biology, and astronomy at The Master's University. His professional memberships include the Creation Research Society, the American Fisheries Association, Southern California Academy of Sciences, Yellowstone Association, and Au Sable Institute of Environmental Studies.

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