

WORKSHEET
CHEMISTRY FOR ENGINEERING STUDENTS, LECTURE 1
 Molecular formula of hydrocarbon or a mixture of hydrocarbons

ANSWERS KEY

QUESTION 1. Calculate the empirical formula of hydrocarbon and its molecular formula.

1. A hydrocarbon of 0.100 g is completely burnt to give 0.345 g carbon dioxide and 0.070 g water. Calculate the empirical formula of the hydrocarbon.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the relative atomic weights of elements as they appear in the balanced eqn
 $12x + 1y + 32z = 44x + 18y$
 - (c) calculate mass of the hydrocarbon from the given data
 0.100 g

2. A 0.100 g hydrocarbon is burnt to produce 0.345 g carbon dioxide and 0.070 g water. Calculate the empirical formula of the hydrocarbon.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the relative atomic weights of elements as they appear in the balanced eqn
 $12x + 1y + 32z = 44x + 18y$
 - (c) calculate mass of the hydrocarbon from the given data
 0.100 g

3. A 0.100 g hydrocarbon is burnt to produce 0.345 g carbon dioxide and 0.070 g water. Calculate the empirical formula of the hydrocarbon.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate mass of the hydrocarbon from the given data
 0.100 g
 - (c) calculate the relative atomic weights of elements as they appear in the balanced eqn
 $12x + 1y + 32z = 44x + 18y$

QUESTION 2. Calculate the empirical formula of the hydrocarbon.

1. A 0.100 g hydrocarbon is burnt to produce 0.345 g carbon dioxide and 0.070 g water. Calculate the empirical formula of the hydrocarbon.
 - (a) write the balanced chemical equation for the reaction
 $C_xH_y + O_2 \rightarrow CO_2 + H_2O$
 - (b) calculate the mass of the hydrocarbon from the given data
 0.100 g

Chem 21 Lab Answers

J Spring



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