

20 Points

Name: Phyllis A. Smith LAB NUMBER: 200

Suppose that you had the correct equation for the following composition for fresh steam. What boiling point would you expect?

CO₂ 18.0%

H₂O

8%

N₂

140°C
boiling

100°C boiling

100°C

100°C
boiling

1. If you had these same conditions at a 1.00 bar total pressure, what boiling point would you expect? (Assume that water is pure at 1.00 bar total pressure.)

$$\frac{18.0 \text{ g CO}_2}{100 \text{ g total}} \left| \frac{44 \text{ g/mol}}{44 \text{ g/mol}} \right| \frac{8 \text{ g N}_2}{100 \text{ g total}} \left| \frac{28 \text{ g/mol}}{28 \text{ g/mol}} \right| \frac{84 \text{ g H}_2\text{O}}{100 \text{ g total}} \left| \frac{18 \text{ g/mol}}{18 \text{ g/mol}} \right| = 100.0^\circ\text{C}$$

2. If you were in the mixture and wanted to do some experiments at 1.00 bar total pressure, what boiling point would you expect?

$$M_1 = \frac{18 \text{ g CO}_2}{100 \text{ g total}} \left| \frac{44 \text{ g/mol}}{44 \text{ g/mol}} \right| \frac{8 \text{ g N}_2}{100 \text{ g total}} \left| \frac{28 \text{ g/mol}}{28 \text{ g/mol}} \right| \frac{84 \text{ g H}_2\text{O}}{100 \text{ g total}} \left| \frac{18 \text{ g/mol}}{18 \text{ g/mol}} \right| = 27.17 \text{ mol } H_2O$$

$$M_2 = \frac{18 \text{ g CO}_2}{80 \text{ g H}_2O} \left| \frac{44 \text{ g/mol}}{44 \text{ g/mol}} \right| \frac{8 \text{ g N}_2}{80 \text{ g H}_2O} \left| \frac{28 \text{ g/mol}}{28 \text{ g/mol}} \right| = 27.17 \text{ mol } H_2O$$

100°C boiling

100°C boiling

100°C boiling

Chem 110 Lab Questions And Answers

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1. Understanding the eBook Chem 110 Lab Questions And Answers
 - The Rise of Digital Reading Chem 110 Lab Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 110 Lab Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 110 Lab Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 110 Lab Questions And Answers
 - Personalized Recommendations
 - Chem 110 Lab Questions And Answers User Reviews and Ratings
 - Chem 110 Lab Questions And Answers and Bestseller Lists
5. Accessing Chem 110 Lab Questions And Answers Free and Paid eBooks
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 - Chem 110 Lab Questions And Answers eBook Subscription Services
 - Chem 110 Lab Questions And Answers Budget-Friendly Options
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 - Chem 110 Lab Questions And Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem 110 Lab Questions And Answers
 - Highlighting and Note-Taking Chem 110 Lab Questions And Answers
 - Interactive Elements Chem 110 Lab Questions And Answers
 8. Staying Engaged with Chem 110 Lab Questions And Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem 110 Lab Questions And Answers
 9. Balancing eBooks and Physical Books Chem 110 Lab Questions And Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 110 Lab Questions And Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chem 110 Lab Questions And Answers
 - Setting Reading Goals Chem 110 Lab Questions And Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chem 110 Lab Questions And Answers
 - Fact-Checking eBook Content of Chem 110 Lab Questions And Answers
 - Distinguishing Credible Sources
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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