

balance nuclear reaction equations

aluminum-26 undergoes positron emission
 potassium-40 undergoes beta decay (electron)
 uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.
 Write the electron configuration for Cr.
 Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride SF_6
- b. Lithium nitride Li_3N
- c. Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride SnCl_2
- e. Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- f. Mercury (I) chloride Hg_2Cl_2
- g. Potassium bromate KBrO_3
- h. Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- i. Chloric acid HClO_3
- j. Sulfurous acid H_2SO_3

Name each of the following compounds:

- a. CuSO_4 Copper (II) sulfate
- b. AlF_3 aluminum fluoride
- c. HI hydrogen iodide
- d. NO nitrogen monoxide
- e. $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- f. HNO_3 nitric acid
- g. HNO_2 nitrous acid
- h. NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?



Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?



Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.



Chapter 7 Modern Chemistry Answers

Camilla Rothe



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Table of Contents Chapter 7 Modern Chemistry Answers

1. Understanding the eBook Chapter 7 Modern Chemistry Answers
 - The Rise of Digital Reading Chapter 7 Modern Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 7 Modern Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 7 Modern Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 7 Modern Chemistry Answers

- Personalized Recommendations
 - Chapter 7 Modern Chemistry Answers User Reviews and Ratings
 - Chapter 7 Modern Chemistry Answers and Bestseller Lists
5. Accessing Chapter 7 Modern Chemistry Answers Free and Paid eBooks
- Chapter 7 Modern Chemistry Answers Public Domain eBooks
 - Chapter 7 Modern Chemistry Answers eBook Subscription Services
 - Chapter 7 Modern Chemistry Answers Budget-Friendly Options
6. Navigating Chapter 7 Modern Chemistry Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Chapter 7 Modern Chemistry Answers Compatibility with Devices
 - Chapter 7 Modern Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chapter 7 Modern Chemistry Answers
 - Highlighting and Note-Taking Chapter 7 Modern Chemistry Answers
 - Interactive Elements Chapter 7 Modern Chemistry Answers
8. Staying Engaged with Chapter 7 Modern Chemistry Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 7 Modern Chemistry Answers
9. Balancing eBooks and Physical Books Chapter 7 Modern Chemistry Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 7 Modern Chemistry Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 7 Modern Chemistry Answers
- Setting Reading Goals Chapter 7 Modern Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 7 Modern Chemistry Answers

- Fact-Checking eBook Content of Chapter 7 Modern Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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