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**Commonwealth Shipping Committee**

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web in order to balance a chemical equation the quantities of each type of element and polyatomic ion that are present in the

reactants and the products of the reaction must be determined because polyatomic ions contain multiple atoms and generally react as indivisible units any polyatomic ion that is present in both a reactant and a product in

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web a balanced chemical equation often may be derived from a qualitative description of some chemical reaction by a fairly simple approach known as balancing by inspection consider as an example the decomposition of water to yield molecular hydrogen and oxygen

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web feb 19 2022 in a balanced chemical equation both the numbers of each type of atom and the total charge are the same

on both sides equations ref 3 1 1 and ref 3 1 2 are balanced chemical equations what is different on each side of the equation is how the atoms are arranged to make molecules or ions

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web mar 14 2023 the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

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web balancing equations the chemical equation described in section 4 1 is balanced meaning that equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides this is a requirement the equation must satisfy to be consistent with the law of conservation of matter

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web because of the law of conservation of mass balancing chemical equations gives a more accurate representation on what is happening in a reaction the accuracy of the ratios of moles is important in calculating the theoretical yield and how much stuff is reacting mass wise stoichiometry

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web a key step in balancing chemical equations is correctly identifying the formulas of the reactants and products for example consider the reaction between calcium oxide  $\text{CaO}$  and  $\text{H}_2\text{O}$  to form aqueous calcium hydroxide  $\text{Ca(OH)}_2$  is it possible to balance the equation if you incorrectly identify the product as  $\text{CaOH}$  and if so what is the equation

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web dec 2 2020 a balanced chemical equation tells you the amounts of reactants and products needed to satisfy the law of conservation of mass basically this means there are the same numbers of each type of atoms on the left side of the equation as there are on the right side of the equation

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web balance the following chemical equation  $\text{mg} + \text{oh} + 2 \text{hcl} \rightarrow \text{mgcl}_2 + \text{h}_2\text{o}$  note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem do 4 problems

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