

Science

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(Chapter - 4) (Structure of the Atom) (Exercises)

(Class - IX)

Question 1:

Compare the properties of electrons, protons and neutrons.

Answer 1:

Electron	Proton	Neutron
(i) Electron are present outside the nucleus of an atom.	(i) Proton are present in the nucleus of an atom.	(i) Neutron are present in the nucleus of an atom.
(ii) Electron are negatively charged	(ii) Proton are positively charged.	(ii) Neutron are neutral.
(iii) The mass of an electron is considered to negligible.	(iii) The mass of a proton is approximately 2000 times as the mass of an electron	(iii) The mass of neutron is nearly equal to the mass of a proton.

Question 2:

What are the limitations of J.J. Thomson's model of the atom?

Answer 2:

According to J.J. Thomson's model of an atom, an atom consists of a positively charged sphere with electrons embedded in it. However, it was later found that the positively charged particles reside at the center of the atom called the nucleus, and the electrons revolve around the nucleus.

Question 3:

What are the limitations of Rutherford's model of the atom?

Answer 3:

According to Rutherford's model of an atom, electrons revolve around the nucleus in fixed orbits. But, an electron revolving in circular orbits will not be stable because during revolution, it will experience acceleration. Due to acceleration, the electrons will lose energy in the form of radiation and fall into the nucleus. In such a case, the atom would be highly unstable and collapse.

Question 4:

Describe Bohr's model of the atom.

Answer 4:

Bohr's model of the atom Niels Bohr proposed the following postulates regarding the model of the atom.

- > Only certain orbits known as discrete orbits of electrons are allowed inside the atom.
- > While revolving in these discrete orbits, the electrons do not radiate energy. These discrete orbits or shells are shown in the following diagram.



The first orbit (i.e., for $n = 1$) is represented by letter K. Similarly, for $n = 2$, it is L - shell, for $n = 3$, it is M - shell and for $n = 4$, it is N - shell. These orbits or shells are also called energy levels.

Investment Science Chapter 4

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Pflug**

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Science and Technology. Panel on Chemical Industry,1973 Introduction to Investments George A. Christy,John C. Clendenin,1982 **Moody's Manual of Investments: American and Foreign** ,1924 National Science Council Review ,2007 *Science and Innovation* Aldo Geuna,Ammon J. Salter,William Edward Steinmueller,2003 The chapters in this book issue a series of challenges to the next generation of science and technology policy *Science* John Michels (Journalist),1889 The Saturday Review of Politics, Literature, Science and Art ,1877 **The Training of Teachers for the Country Schools of Wisconsin** Annie Reynolds,1917 **Moody's Analyses of Railroad Investments** John Sherman Porter,1925 American government securities 1928 53 in 5 annual vols v 1 Railroad securities 1952 53 Transportation v 2 Industrial securities v 3 Public utility securities v 4 Government securities 1928 54 v 5 Banks insurance companies investment trusts real estate finance and credit companies 1928 54 An Introduction to Management Science David Ray Anderson,John S. Loucks,2000 **British Journal of Dental Science** ,1930 *Annals of the American Academy of Political and Social Science* ,1920

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