

# 9

## Thermodynamics

1. Which of the following statements is correct for the spontaneous absorption of a gas? [2014]
- (a)  $\Delta S$  is negative and therefore,  $\Delta H$  should be highly positive  
(b)  $\Delta S$  is negative and therefore,  $\Delta H$  should be highly negative  
(c)  $\Delta S$  is positive and therefore,  $\Delta H$  should be negative  
(d)  $\Delta S$  is positive and therefore,  $\Delta H$  should also be highly positive
2. For the reaction,  $X_2O_4(l) \longrightarrow 2XO_2(g)$ ,  $\Delta U = 2.1 \text{ kcal}$ ,  $\Delta S = 20 \text{ cal K}^{-1}$  at 300 K. Hence,  $\Delta G$  is [2014]
- (a) 2.7 kcal (b) - 2.7 kcal  
(c) 9.3 kcal (d) - 9.3 kcal
3. For a given exothermic reaction,  $K_p$  and  $K_p'$  are the equilibrium constants at temperatures  $T_1$  and  $T_2$ , respectively. Assuming that heat of reaction is constant in temperature range between  $T_1$  and  $T_2$ , it is readily observed that [2014]
- (a)  $K_p > K_p'$   
(b)  $K_p < K_p'$   
(c)  $K_p = K_p'$   
(d)  $K_p = \frac{1}{K_p'}$
4. A reaction having equal energies of activation for forward and reverse reactions has [2013]
- (a)  $\Delta S = 0$  (b)  $\Delta G = 0$   
(c)  $\Delta H = 0$  (d)  $\Delta H = \Delta G = \Delta S = 0$
5. In which of the following reactions, standard reaction entropy changes ( $\Delta S^\circ$ ) is positive and standard Gibbs energy change ( $\Delta G^\circ$ ) decreases sharply with increasing temperature? [2012]
- (a)  $C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow CO(g)$   
(b)  $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$   
(c)  $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$   
(d)  $\frac{1}{2}C(\text{graphite}) + \frac{1}{2}O_2(g) \longrightarrow \frac{1}{2}CO_2(g)$
6. The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is [2012]
- (a) 10.52 cal/mol K (b) 21.04 cal/mol K  
(c) 5.260 cal/mol K (d) 0.526 cal/mol K
7. Standard enthalpy of vaporisation  $\Delta_{\text{vap}}H^\circ$  for water at 100°C is 40.66 kJ mol<sup>-1</sup>. The internal energy of vaporisation of water at 100°C (in kJ mol<sup>-1</sup>) is (assume water vapour to behave like an ideal gas). [2012]
- (a) + 37.56 (b) - 43.76  
(c) + 43.76 (d) + 40.66
8. If the enthalpy change for the transition of liquid water to steam is 30 kJ mol<sup>-1</sup> at 27°C, the entropy change for the process would be [2011]
- (a) 1.0 J mol<sup>-1</sup>K<sup>-1</sup>  
(b) 0.1 J mol<sup>-1</sup>K<sup>-1</sup>  
(c) 100 J mol<sup>-1</sup>K<sup>-1</sup>  
(d) 10 J mol<sup>-1</sup>K<sup>-1</sup>

# Engineering Thermodynamics Important Questions Answers

**Zhe Xu**



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