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1

Chemistry Lab Experiment

Determination of the Molar Volume of a Gas

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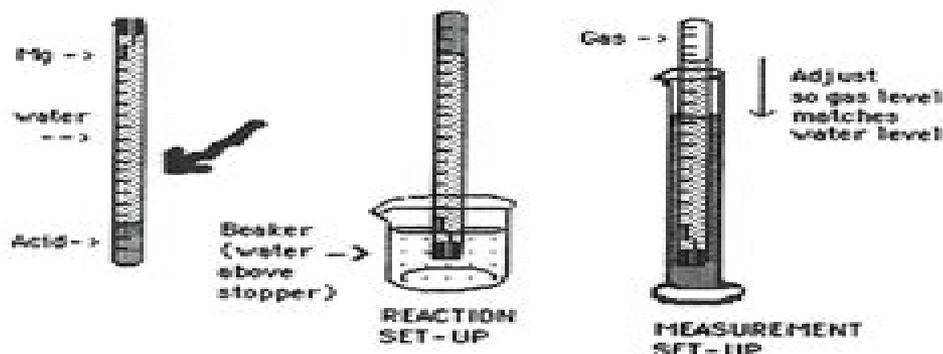
Name:

Lab Partner(s):

Purpose:

- To experimentally determine the molar volume of hydrogen gas and to compare it with the $\frac{22.4\text{-L}}{\text{mole}}$ theoretical STP value.
- To relate gas law concepts learned in MA Frameworks Topic 6 (e.g., the Combined Gas Law and Dalton's Law of Partial Pressures) to previously learned concepts involving balancing chemical equations and stoichiometric relationships.

Introduction:



The **water displacement method** can be used to capture the hydrogen gas liberated when a strip of magnesium metal reacts with a 3.0-*M* solution of hydrochloric acid. The total gas volume of H₂O vapor and H₂ adjusted to the room's barometric pressure can be determined from the **eudiometer** (graduated gas-measuring tube) by partially immersing the tube into a large container of water until the liquid level inside the tube matches the water level in the large container.

According to **Dalton's Law of Partial Pressures**, the total gas pressure in the eudiometer is the sum of the pressure of the hydrogen gas collected plus the **water vapor pressure**. Symbolically, $P_{\text{total}} = P_{\text{H}_2} + P_{\text{H}_2\text{O}}$. Subtracting the water vapor pressure from the total pressure gives the pressure of the dry hydrogen gas.

Determining Molar Volume Gas Post Lab Answers

Albert A Gayle



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Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

Elements of Experimental Chemistry René Lahaie, Luc Papillon, Pierre Valiquette, 1978

Experiments in General Chemistry Kenneth W. Whitten, Carl B. Bishop, Kenneth D. Gailey, 1992

Working with Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well

Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations

Energy Research Abstracts, 1978

Scientific and Technical Aerospace Reports, 1980 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database

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