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**Julia Schneider**



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web the orbital diagram of helium is  $1s^2$  the spin of an electron is specified by a fourth quantum number called the the two possible values of  $m_s$  are spin up and spin down the three principles that are needed to specify orbital occupancy for the electrons are electrons and they must have pauli exclusion principle no orbital may have more than

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web the electron configuration and the orbital diagram are following hydrogen is the noble gas helium which has an atomic number of 2 the helium atom contains two protons and two electrons the first electron has the same four quantum numbers as the hydrogen atom electron  $n=1, l=0, m_l=0, m_s=1/2$

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the various orbitals available to electrons this table is easy to remember and it makes it possible to generate the electron configuration table for

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web chemistry chemistry questions and answers part awrite the full electron configuration for the following si3 part bwrite the noble gas abbreviated electron configuration for the following nb4 part cdraw the orbital diagram for the following excited ca2 determine the number of unpaired electrons from your previous answer