

## 12

## STOICHIOMETRY

## Practice Problems

In your notebook, solve the following problems.

## SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate,  $\text{KClO}_3$ . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from  $3.8 \times 10^{22}$  molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

## SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(III) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 15.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. (Write a balanced equation first.)

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of nitrogen dioxide. Assume all gases are at the same conditions of temperature and pressure.



# Pearson Chemistry Chapter 12 Stoichiometry

**Emil J. Margolis**



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**Teaching Science for Understanding** James Joseph Gallagher, 2007 Offers middle and high school science teachers practical advice on how they can teach their students key concepts while building their understanding of the subject through various levels of learning activities

**Prentice Hall Chemistry** Antony C. Wilbraham, 2008

**The Group 13 Metals Aluminium, Gallium, Indium and Thallium** Simon Aldridge, Anthony J. Downs, 2011-04-11 The last two decades have seen a renaissance in interest in the chemistry of the main group elements In particular research on the metals of group 13 aluminium gallium indium and thallium has led to the synthesis and isolation of some very novel and unusual molecules with implications for organometallic synthesis new materials development and with biological medical and environmental relevance *The Group 13 Metals Aluminium Gallium Indium and Thallium* aims to cover new facts developments and applications in the context of more general patterns of physical and chemical behaviour Particular attention is paid to the main growth areas including the chemistry of lower formal oxidation states cluster chemistry the investigation of solid oxides and hydroxides advances in the formation of III V and related compounds the biological significance of Group 13 metal complexes and the growing importance of the metals and their compounds in the mediation of organic reactions Chapters cover general features of the group 13 elements group 13 metals in the 3 oxidation state simple inorganic compounds formal oxidation state 3 organometallic chemistry formal oxidation state 2 metal metal bonded vs mononuclear derivatives group 13 metals in the 1 oxidation state mixed or intermediate valence group 13 metal compounds aluminium and gallium clusters metalloid clusters and their relation to the bulk phases to naked clusters and to nanoscaled materials simple and mixed metal oxides and hydroxides solids with extended structures of different dimensionalities and porosities coordination and solution

chemistry of the metals biological medical and environmental relevance III V and related semiconductor materials group 13 metal mediated organic reactions The Group 13 Metals Aluminium Gallium Indium and Thallium provides a detailed wide ranging and up to date review of the chemistry of this important group of metals It will find a place on the bookshelves of practitioners researchers and students working in inorganic organometallic and materials chemistry

**Inorganic Fluorine Chemistry** Joseph S. Thrasher,1994 Offers an up to date account of fundamental and applied research and points to future directions in inorganic fluorine chemistry Reviews the major advances in fluorine containing substituent groups Provides a critical examination of coordination numbers greater than six among fluorides and oxofluorides Discusses attempts to control chemical reactivity and stability through the use of more nucleophilic sources of fluoride or conversely more weakly coordinating anions Features chapters on new fluorine containing ligands in organometallic transition metal solid state lanthanide and actinide chemistry

*Elements of Inorganic Structural Chemistry* Erwin Parthé,1996

*Porous Silicon: From Formation to Application: Biomedical and Sensor Applications, Volume Two* Ghenadii Korotcenkov,2016-01-05 Porous silicon is rapidly attracting increasing interest from various fields including optoelectronics microelectronics photonics medicine chemistry and biosensing This nanostructured and biodegradable material has a range of unique properties that make it ideal for many applications For example the pores and surface chemistry of the mater

Fundamentals of Water Treatment Unit Processes David Hendricks,2016-04-19 Carefully designed to balance coverage of theoretical and practical principles Fundamentals of Water Treatment Unit Processes delineates the principles that support practice using the unit processes approach as the organizing concept The author covers principles common to any kind of water treatment for example drinking water municipal wastew

**The Journal of Physics and Chemistry of Solids** ,1971

**Students Study Guide** Fredeen,Pearson,2003-07

**The Software Encyclopedia** ,1988

**City Record** Boston (Mass.),1918

**Chemical Engineering** John Metcalfe Coulson,1971 *Chemical Engineering: Chemical reactor design, biochemical reaction engineering including computational techniques and control* John Metcalfe Coulson,John Francis Richardson,1979

**Chemical Engineering: Richardson, J. F. and Peacock, D. G. Chemical reactor design, biochemical reaction engineering including computational techniques and control** John Metcalfe Coulson,1968

**Zeitschrift Für Kristallographie** ,2000

**Water Treatment Principles and Design** James M. Montgomery, Consulting Engineers,1985-09-03 Drawing on the vast experience of the most respected firm in the industry Water Treatment Principles and Design is the first major reference on the science of water treatment in several decades It covers both the practical and theoretical aspects of water quality analysis treatment plant operation and facility design and provides detailed descriptions of processes such as coagulation and flocculation sedimentation filtration ion exchange and adsorption In addition it offers one of the most extensive discussions ever published on design criteria including component description and organization aeration equipment upflow clarifiers disinfection and materials

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web procedure step 1 set up a ramp with the angle of the incline at about 10 to the horizontal as shown this can be an estimate the exact angle is not important step 2 divide the

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web background measurement of the motion of a freely falling object is difficult because the speed is constantly increasing by approximately 10 m/s every second the distance

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web speeds in the table also record the average rolling time it took for the ball to travel each distance from the release point from the information recorded in data table c make a

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