

Solutions for Gravimetric Analysis Questions

Check for Understanding 4.1

1. Determine the solubility of AgCl using K_{sp} for AgCl and a table of initial and equilibrium concentration terms.

	$\text{AgCl}(s)$	\rightleftharpoons	$\text{Ag}^+(\text{aq})$	$+$	$\text{Cl}^-(\text{aq})$	
init			0.0060		0	
equil			$0.0060 + x$		x	where x = increase in $[\text{Cl}^-]$ = solubility of AgCl

at equil, $K_{sp} = 1.8 \times 10^{-10} = [\text{Ag}^+][\text{Cl}^-] = (0.0060 + x)(x) = 0.0060x \rightarrow x = 3.0_0 \times 10^{-8} \text{ M}$

2. $200. \text{ mL} \times \frac{3.0_0 \times 10^{-8} \text{ mol AgCl}}{1000 \text{ mL}} \times \frac{143.32 \text{ g AgCl}}{\text{mol AgCl}} = 8.6 \times 10^{-7} \text{ g AgCl}$

Check for Understanding 4.2

1. Often the solubility of a slightly soluble salt is increased in acidic solution because the anion combines with H^+ to form a weak acid. Removal of the anion shifts the equilibrium toward more dissolution of the salt. Since Cl would form a strong acid (HCl) that is ~100% ionized in solution, it is not removed from solution. Therefore, AgCl is not more soluble in acidic solution and the presence of HNO_3 does not increase its solubility.

Check for Understanding 4.3

1. First, determine the solubility of AgCl using K_{sp} for AgCl and a table of initial and equilibrium concentration terms.

	$\text{AgCl}(s)$	\rightleftharpoons	$\text{Ag}^+(\text{aq})$	$+$	$\text{Cl}^-(\text{aq})$	
init			0		0	
equil			x		x	

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Chandresh Agrawal, Nandini Books

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