

# Physical Chemistry Formula Revision

## Basic Concepts

$$\text{Average atomic mass} = \frac{a_1x_1 + a_2x_2 + a_3x_3 + \dots}{100}$$

$a_1, a_2, a_3, \dots$  - atomic mass of isotopes

$x_1, x_2, x_3, \dots$  - percentage abundance

## Mole concept

$$\text{no. of moles, } n = \frac{\text{given mass}}{\text{molar mass}} = \frac{W}{M}$$

$$n = \frac{\text{no. of particles}}{\text{Avogadro's No.}} = \frac{N}{6.022 \times 10^{23}}$$

$$n = \frac{\text{vol. of gas}}{\text{vol. of gas at STP}} = \frac{V}{22.4 \text{ L}}$$

## Concentration Terms

$$\text{Percentage, } \% = \frac{W_2 \text{ or } V_2}{W_1 \text{ or } V_1} \times 100$$

$$\text{Molarity, } M = \frac{n_2}{V(l)} = \frac{W_2 \times 1000}{M_2 \times V(\text{ml})}$$

$$\text{Normality, } N = \frac{n_2 \times 1000}{V(l)} = \frac{W_2 \times 1000}{E_2 \times V(\text{ml})}$$

$\hookrightarrow \frac{M_2}{n_f}$

$$\text{parts per million, ppm} = \frac{W_2}{W_1} \times 10^6$$

$$\text{Molality, } m = \frac{n_2}{M_2(\text{kg})} = \frac{W_2 \times 1000}{M_2 \times W_1(\text{g})}$$

$$\text{Mole fraction, } X_2 = \frac{n_2}{n_1 + n_2 + n_3 + \dots}$$

## Stoichiometry



$$\frac{\text{moles of } A}{a} = \frac{\text{moles of } B}{b} = \frac{\text{moles of } C}{c}$$

Limiting  $R_2$ : given amt.  $\rightarrow$  lowest stoichiometric coeff.

## Atomic Structure

### Planck's Quantum Theory

$$E = \frac{hc}{\lambda} \quad E(\text{eV}) = \frac{12400}{\lambda(\text{\AA})}$$

## Photo Electric Effect



$$K_{\text{max}} = h\nu - \phi + K_{\text{max}}$$

$$e = h\nu_0 + \frac{m v_{\text{max}}^2}{2}$$



## De Broglie's Hypothesis

$$m\lambda = \frac{h\nu}{\lambda} \rightarrow \lambda = \frac{h}{mv} \quad \lambda = \frac{h}{\sqrt{2mE}}$$

Bohr's Atomic model

$$mvr = \frac{nh}{2\pi} \quad v = 0.53 \times \frac{n^2}{r} \text{ \AA}^{-1}$$

$$r = 0.529 \times 10^{-10} \frac{n^2}{Z} \text{ m}$$

$$E_n = -13.6 \times \frac{Z^2}{n^2} \text{ eV} \quad T_n = \frac{h}{mv} = \frac{h}{2\pi mvr} = \frac{h}{2\pi n}$$

$$\frac{1}{\lambda} = R_H Z^2 \left[ \frac{1}{n_1^2} - \frac{1}{n_2^2} \right] \quad R_H = 1.097 \times 10^7 \text{ m}^{-1}$$

max no. of spectral lines when  $e^-$  returns from  $n_2 = n_1$   $(\frac{n_2(n_2-1)}{2} - \frac{n_1(n_1-1)}{2})$

Heisenberg's uncertainty principle:  $\Delta x \Delta p \geq \frac{h}{4\pi}$

## Quantum Model

$$\text{total nodes} = n - 1$$

$$\text{angular nodes} = l$$

$$\text{radial nodes} = n - l - 1$$

## Angular momentum

$$\text{orbital} = \sqrt{l(l+1)} \frac{h}{2\pi}$$

$$\text{spin} = \sqrt{s(s+1)} \frac{h}{2\pi}$$

# Dictionary Of Physical Chemistry

**Mark Steyvers**



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