

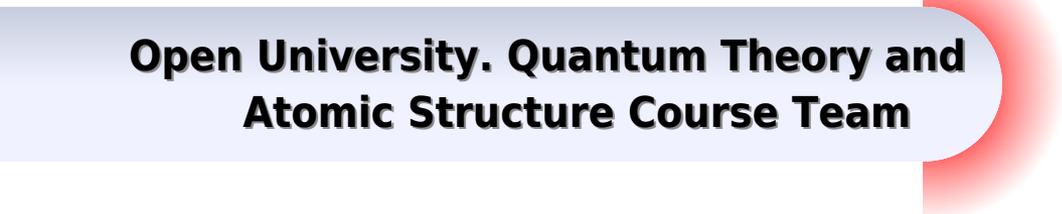
Complete the following:

Element	Electron configuration	Orbital Notation	Noble Gas Notation	Lewis Electron Dot Notation
1. Chromium				
2. Tin	Not needed	Not Needed		
3. Aluminum				
4. Helium				
5. Gold	Not Needed	Not Needed		

5. What are the elements that have the following electron configurations?
- a)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 5d^1 4f^4$  \_\_\_\_\_
- b)  $1s^2 2s^2 2p^4$  \_\_\_\_\_
- c)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^7$  \_\_\_\_\_
- d)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6$  \_\_\_\_\_
6. How many orbitals does each of the following subshells have?
- a) s \_\_\_\_\_      b) d \_\_\_\_\_      c) f \_\_\_\_\_      d) p \_\_\_\_\_
7. How many electrons does each of the following subshells have?
- a) s \_\_\_\_\_      b) d \_\_\_\_\_      c) f \_\_\_\_\_      d) p \_\_\_\_\_
8. How many electrons can occupy each of the following energy levels?
- a) l \_\_\_\_\_      b) 6 \_\_\_\_\_      c) 3 \_\_\_\_\_      d) 4 \_\_\_\_\_
9. Describe the property of the electron that each of the four Quantum Numbers exhibit.
- n - Principal Quantum Number
- l - Angular Momentum Quantum Number
- m<sub>l</sub> - Magnetic Quantum Number

# Quantum Theory And The Atom Worksheet Answers

**Open University. Quantum Theory and  
Atomic Structure Course Team**



## Quantum Theory And The Atom Worksheet Answers:

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and some fundamental problems involved in the description of atomic and sub atomic systems are examined An introduction to quantum theory with some key examples is also provided The emphasis throughout is on concepts References to experiments are made to test and support theories and introduce new information Mathematical complexity has been avoided as much as possible and problems of both a formal analytical and numerical kind are

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