

# 2025 MORTGAGE RATE FORECAST



30-Year Fixed  
Mortgage Rate

2017

2019

2021

2023

# Mortgage Rates This Month

**JL Elias**

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular glow behind it.

## **Mortgage Rates This Month:**

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web oxidation reduction redox reactions oxidation  $\text{Cu} + \text{Cu}^{2+} + 2\text{e}^-$  reduction  $\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$  l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients  $\text{Cu} + \text{MnO}_4^- + 4\text{H}^+$  is

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web 1 practice problem balance the redox reaction in basic solution  $\text{NO}_2 + \text{Al} + \text{OH}^- \rightarrow \text{AlO}_2^- + \text{NO}_2^- + \text{H}_2\text{O}$  2 practice problem balance the following reduction oxidation reaction in basic solution what is the coefficient in front of  $\text{OH}^-$   $\text{Br}_2 + \text{I}^- \rightarrow \text{Br}^- + \text{I}_2$  previous topic next topic

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web google classroom for the following redox equation  $\text{MnO}_4^- + \text{SO}_2 + 2\text{H}^+ \rightarrow \text{Mn}^{2+} + \text{HSO}_4^- + \text{H}_2\text{O}$  what will be the stoichiometric coefficient of  $\text{HSO}_4^-$  if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- \rightarrow 2\text{H}_2\text{O} + \text{H}_2\text{O}_2 + 2\text{e}^-$  make electrons equal

convert to basic solution  $4\text{CN}^- + 2\text{Au} + \text{CN}^- + 2\text{e}^-$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 + \text{H}_2\text{O}_2 + 2\text{OH}^-$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add OH<sup>-</sup> ions or the H<sub>2</sub>O pair to fully balance the equation in this video we ll walk through this process for the reaction between ClO<sup>-</sup> and Cr(OH)<sub>4</sub> in basic solution created by jay

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation Ni(s) → Ni<sup>2+</sup>(aq) + 2e<sup>-</sup> now that the oxidation half reaction is balanced it tells us that two electrons are produced for

every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  2 equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web when balancing equations for redox reactions occurring in acidic solution it is often necessary to add  $\text{H}^+$  ions or the  $\text{H}_2\text{O}$  pair to fully balance the equation in this video we'll walk through this process for the reaction between dichromate  $\text{Cr}_2\text{O}_7^{2-}$  and chloride  $\text{Cl}^-$  ions in acidic solution created by jay

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web balance each half reaction in basic solution 4 5 6 7  $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 + \text{SO}_4^{2-} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method 8  $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow \text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te}$  10  $\text{ReO}_4^- + \text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding  $\text{H}^+$   $\text{H}_2\text{O}$  and  $\text{e}^-$  in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 + 2\text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} \rightarrow 2\text{NO} + \text{C}_6\text{H}_8\text{O}_6 + \text{NO}_2 + \text{C}_6\text{H}_8\text{O}_6 + \text{NO}_2 + 2\text{H}_2\text{O}$  choose 1 answer c  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  a  $\text{C}_6\text{H}_8\text{O}_6$  and  $\text{NO}_2$  b  $\text{NO}_2$  and  $\text{C}_6\text{H}_8\text{O}_6$  h and  $\text{NO}_2$  c h and