

CHEMISTRY



Atomic Structure

Atoms are made up of three subatomic particles: electrons, protons and neutrons. Electrons are negatively charged, protons are positively charged and neutrons are neutral. The mass of an electron is very small compared to that of a proton and neutron. The mass of a proton is approximately 1.67 x 10⁻²⁷ kg and the mass of a neutron is approximately 1.67 x 10⁻²⁷ kg. The mass of an electron is approximately 9.1 x 10⁻³¹ kg.

Atomic Number

The atomic number of an element is the number of protons present in the nucleus of an atom of that element. It is denoted by the symbol Z.

Atomic Mass

The atomic mass of an element is the sum of the number of protons and neutrons present in the nucleus of an atom of that element. It is denoted by the symbol A.

Isotopes

Isotopes are atoms of the same element having the same atomic number but different mass numbers. They have the same chemical properties but different physical properties.

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Element	Atomic Number (Z)	Atomic Mass (A)
Hydrogen	1	1
Helium	2	4
Lithium	3	7
Boron	5	11
Carbon	6	12
Nitrogen	7	14
Oxygen	8	16
Fluorine	9	19
Neon	10	20
Sodium	11	23
Magnesium	12	24
Aluminum	13	27
Silicon	14	28
Phosphorus	15	31
Sulfur	16	32
Chlorine	17	35.5
Argon	18	40
Potassium	19	39
Calcium	20	40

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