

NAME \_\_\_\_\_ DATE \_\_\_\_\_ CLASS \_\_\_\_\_

**CHAPTER 12** **STUDY GUIDE FOR CONTENT MASTERY**

### Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of  $\text{O}_2$ ,  $\text{NO}$ , and  $\text{NO}_2$  are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of $\text{O}_2$	Amount of $\text{NO}$	Amount of $\text{NO}_2$	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
2 molecules	4 molecules	4 molecules	NO	2 molecules $\text{O}_2$
2 molecules	4 molecules	1, 4 molecules	1, $\text{O}_2$	3, 4 molecules NO
1.00 mol	2.00 mol	3, 2.00 mol	3, none	3, none
4.00 mol	6.00 mol	2, 4.00 mol	3, NO	3, 2.00 mol $\text{O}_2$
6.00 mol	1.00 mol	10, 7.00 mol	11, NO	10, 1.00 mol $\text{O}_2$
1.00 mol	3.00 mol	13, 2.00 mol	14, $\text{O}_2$	13, 2.00 mol NO
0.200 mol	0.200 mol	16, 0.200 mol	17, NO	16, 0.400 mol $\text{O}_2$
33.00 g	66.02 g	19, 92.02 g	20, none	21, none
64.00 g	60.00 g	22, 66.02 g	22, $\text{O}_2$	24, 30.12 g NO
16.00 g	30.00 g	25, 28.76 g	26, $\text{O}_2$	27, 1.24 g NO

balanced equation mole ratio = 2 mol NO/1 mol  $\text{O}_2$

16.00 g  $\text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

20.00 g NO  $\times 1 \text{ mol NO} / 30.01 \text{ g NO} = 0.6664 \text{ mol NO}$

actual mole ratio = 0.6664 mol NO/0.5125 mol  $\text{O}_2 = 2.132 \text{ mol NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of NO/ $\text{O}_2$  is larger than the balanced equation mole ratio of NO/ $\text{O}_2$ , there is an excess of NO;  $\text{O}_2$  is the limiting reactant.

Mass of NO used = 0.5125 mol  $\text{O}_2 \times 2 \text{ mol NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol NO}$

0.4250 mol NO  $\times 30.01 \text{ g NO} / 1 \text{ mol NO} = 12.76 \text{ g NO}$

Mass of  $\text{NO}_2$  produced = 0.4250 mol NO  $\times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol NO} = 19.56 \text{ g } \text{NO}_2$

Excess NO = 20.00 g NO – 12.76 g NO = 7.24 g NO

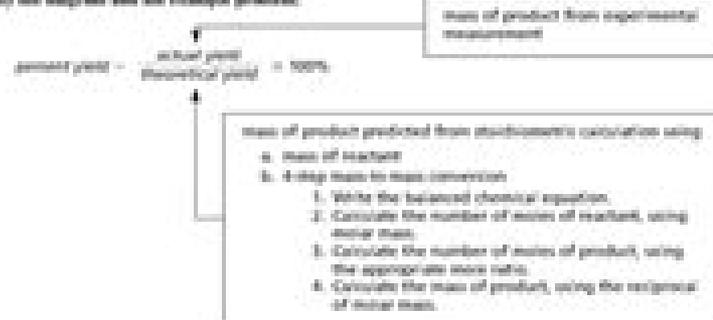
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**CHAPTER 12** **STUDY GUIDE FOR CONTENT MASTERY**

### Section 12.4 Percent Yield

In your textbook, read about the percent of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of yellow oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and  $\text{Ga}_2\text{O}_3$  are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g  $\text{Ga}_2\text{O}_3$
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga  $\times \frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}} = 7.00 \text{ g Ga} \times 1 \text{ mol Ga} / 69.72 \text{ g Ga}$
- number of moles of product: 0.100 mol Ga  $\times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol Ga}} = 0.100 \text{ mol Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol Ga}}$
- theoretical yield: 0.200 mol  $\text{Ga}_2\text{O}_3 \times \frac{187.44 \text{ g } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } (\text{Ga}_2\text{O}_3)} = 37.49 \text{ g } \text{Ga}_2\text{O}_3$  correct
- percent yield:  $\frac{7.42 \text{ g } (\text{Ga}_2\text{O}_3)}{37.49 \text{ g } (\text{Ga}_2\text{O}_3)} \times 100 = 7.42 \text{ g } \text{Ga}_2\text{O}_3 / 37.49 \text{ g } \text{Ga}_2\text{O}_3 \times 100$

# Glencoe Science Chemistry Matter And Change Answer Key Chapter 12

**Rachel Sandford**



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