

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, read about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

octaves	atomic mass	atomic number	nine
elements	properties	Henry Moseley	eight
protons	periodic law	Dmitri Mendeleev	accepted

The table below was developed by John Newlands and is based on a relationship called the law of **(1)** _____. According to this law, the properties of the elements repeated every **(2)** _____ elements. Thus, for example, element two and element **(3)** _____ have similar properties. The law of octaves did not work for all the known elements and was not generally **(4)** _____.

1	2	3	4	5	6	7
H	Li	G	Bo	C	N	O
B	S	10	11	12	13	14
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to **(5)** _____. In his table, the elements were arranged according to increasing **(6)** _____. One important result of this table was that the existence and properties of undiscovered **(7)** _____ could be predicted.

The element in the modern periodic table are arranged according to increasing **(8)** _____, as a result of the work of **(9)** _____. This arrangement is based on number of **(10)** _____ in the nucleus of an atom of the element. The modern form of the periodic table results in the **(11)** _____, which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical **(12)** _____.

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Table of Contents Teaching Transparency The Periodic Table Answers

1. Understanding the eBook Teaching Transparency The Periodic Table Answers
 - The Rise of Digital Reading Teaching Transparency The Periodic Table Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Teaching Transparency The Periodic Table Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Teaching Transparency The Periodic Table Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Teaching Transparency The Periodic Table Answers
 - Personalized Recommendations
 - Teaching Transparency The Periodic Table Answers User Reviews and Ratings
 - Teaching Transparency The Periodic Table Answers and Bestseller Lists
5. Accessing Teaching Transparency The Periodic Table Answers Free and Paid eBooks
 - Teaching Transparency The Periodic Table Answers Public Domain eBooks
 - Teaching Transparency The Periodic Table Answers eBook Subscription Services
 - Teaching Transparency The Periodic Table Answers Budget-Friendly Options

6. Navigating Teaching Transparency The Periodic Table Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Teaching Transparency The Periodic Table Answers Compatibility with Devices
 - Teaching Transparency The Periodic Table Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Teaching Transparency The Periodic Table Answers
 - Highlighting and Note-Taking Teaching Transparency The Periodic Table Answers
 - Interactive Elements Teaching Transparency The Periodic Table Answers
8. Staying Engaged with Teaching Transparency The Periodic Table Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Teaching Transparency The Periodic Table Answers
9. Balancing eBooks and Physical Books Teaching Transparency The Periodic Table Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Teaching Transparency The Periodic Table Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Teaching Transparency The Periodic Table Answers
 - Setting Reading Goals Teaching Transparency The Periodic Table Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Teaching Transparency The Periodic Table Answers
 - Fact-Checking eBook Content of Teaching Transparency The Periodic Table Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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