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partly because so many compounds can act as acids or bases

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a 25 0 ml sample of 0 100 m hydrochloric acid with 0 100 m sodium hydroxide the values of the ph measured after successive
additions of small amounts of naoh are listed in the first column of this table and are graphed in

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until the solution is exactly neutralised i e there is only salt and water it is possible to calculate

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titration of an acid with a base or vice versa the indicator is a substance that can exist in two forms an

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