

Acids, bases and pH

Bronsted-Lowry definitions

Acid: proton donor

Base: proton acceptor

monobasic \rightarrow 1 H^+

dibasic \rightarrow 2 H^+

tribasic \rightarrow 3 H^+

Arrhenius model

Acids dissociate and release H^+ ions in aqueous solution

Alkalis dissociate and release OH^- ions in aqueous solution

H^+ ions are neutralised by OH^- ions to form water $H^+_{(aq)} + OH^-_{(aq)} \rightarrow H_2O$

An acid is a proton donor



• Stronger acid donates proton so other acid acts as a base

• Each side of arrow needs:

① An acid and a base

② Different numbers



$$pH = -\log [H^+]$$

$$[H^+] = 10^{-pH}$$

pH of strong acids

K_a - acid dissociation constant

Higher $K_a \rightarrow$ Stronger acid

Lower $pK_a \rightarrow$ Stronger acid



$$K_a = \frac{[H^+][A^-]}{[HA]}$$

Assume full dissociation so initial $[HA] = [H^+]$ when in solution and $[A^-] = [H^+]$ so $K_a = [H^+]$

$$pK_a = -\log K_a$$

$$K_a = 10^{-pK_a}$$

A Level Chemistry Revision Notes

Karin Nielsen-Saines



A Level Chemistry Revision Notes:

Reviewing **A Level Chemistry Revision Notes**: Unlocking the Spellbinding Force of Linguistics

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