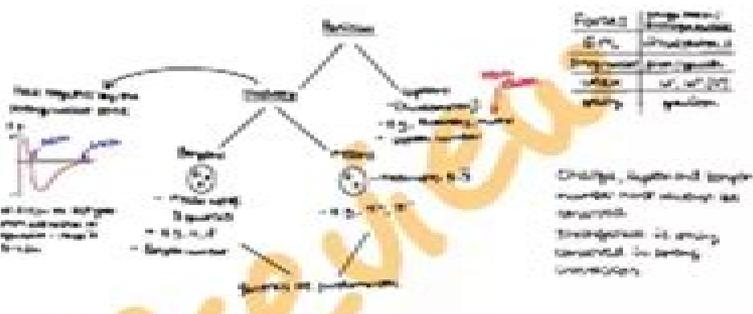


Particles and Quantum



To absorb must be energy... $E_{photon} = h\nu$
 Annihilation - pair of particles and its corresponding antiparticle created and their rest energy is converted into photons



Pair production - when photon turns into a particle and its corresponding antiparticle. The photon must have at least the sum amount of energy which rest energies of particles



Excitation

Exciting means they gain energy through energy to climb their atomic energy levels - ΔE
 Alpha particles - Helium nucleus, ${}^4_2\text{He}$
 - highly ionizing, so it is dangerous
 - relatively penetrating - stopped by a plate of paper or foil on skin
 ${}^{238}_{92}\text{U} \rightarrow {}^{234}_{90}\text{Th} + {}^4_2\text{He}$

Beta - fast moving electron $\rightarrow e^-$
 - medium penetrating and relatively ionizing - stopped by a few cm of aluminium
 - fast moving - ${}^{14}_6\text{C} \rightarrow {}^{14}_7\text{N} + e^- + \bar{\nu}_e$ **neutrino** is almost transparent



Gamma - high energy EM ray or photon emitted from a high energy nucleus $\rightarrow \gamma$
 - highly penetrating - stopped by a few m of concrete

Isotopes - same element - same number of protons
 - different relative atomic mass - different number of neutrons
 ${}^{12}_6\text{C} + {}^{13}_6\text{C}$

Nuclear charge - $\frac{Q}{A} [10^{-19}]$
 - charge to mass ratio

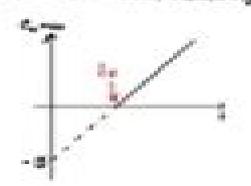
Electron rest - charge of an electron when stationary through a bit of the rest
 - $1.6 \times 10^{-19} \text{ C}$
 - $9.1 \times 10^{-31} \text{ kg}$



Photoelectric effect

Use photons of sufficient energy or absorbed by electrons in the surface of a metal, liberating them

- 1. $E_{photon} = hf$ - energy of photon
- 2. $hf = \phi + E_{kin}$ - energy of electron
- 3. $E_{kin} = hf - \phi$ - kinetic energy of electron
- 4. $E_{photon} = hf$ - energy of photon
- 5. $E_{photon} = hf$ - energy of photon



The photoelectric effect proved the particle nature of light due to the fact that the maximum kinetic energy of electrons E_k is not linearly dependent on the wavelength of light but is linearly dependent on the frequency of light. E_k is not dependent on the intensity of light but is dependent on the number of electrons that are emitted per second.

Stopping potential - the pd needed to stop any electrons from reaching the other side

- 1. $E_{photon} = hf$
- 2. $hf = \phi + E_{kin}$
- 3. $E_{kin} = hf - \phi$
- 4. $E_{photon} = hf$
- 5. $E_{photon} = hf$

De Broglie

The de Broglie wavelength is the wavelength that a particle can have. The wave nature of a particle was proven by diffraction of a crystal. This is because the electron is diffracting which is wave nature.



$\lambda = \frac{h}{p}$
 $\lambda = \frac{h}{mv}$
 $\lambda = \frac{h}{m\gamma v}$
 - λ - de Broglie wavelength
 - p - momentum / mass x velocity

Fluorescent tubes

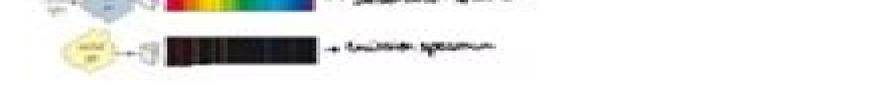


These tubes contain mercury under a high voltage is applied. This excites the gas atoms to which some of the mercury atoms emit photons when they relax. When the gas relaxes it emits photons when the excited mercury atoms emit the higher energy photons when the electrons relax, they emit photons in the UV range. A phosphor coating in the tube absorbs the photons and emit visible light.

Energy levels

Electrons can be excited to higher energy levels. Either by absorbing photons from the light or by the collision. In a gas discharge tube or in a gas discharge cell when the electrons are excited to higher energy levels.

Ionization energy - is enough energy to pull an electron out of an atom, so will completely change the atom.



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