

Specific Heat and Heat Capacity

Name _____ KEY _____ Period _____

1. Calculate the amount of heat needed to increase the temperature of 250 g of water from 20 °C to 46 °C.

$$q = m \times C \times \Delta T$$
$$q = 250 \text{ g} \times 4.18 \text{ J/g}^\circ\text{C} \times 26^\circ\text{C}$$
$$q = 27,170 \text{ J or } 27 \text{ kJ}$$

2. Calculate the specific heat capacity of copper given that 204.75 J of energy raises the temperature of 15 g of copper from 25 °C to 60. °C.

$$q = m \times C \times \Delta T$$
$$C = q/m \times \Delta T$$
$$C = 204.75 \text{ J} / (15 \text{ g} \times 35^\circ\text{C})$$
$$C = 0.39 \text{ J/g}^\circ\text{C}$$

3. 216 J of energy is required to raise the temperature of aluminum from 15 °C to 35 °C. Calculate the mass of aluminum. (Specific heat capacity of aluminum is 0.90 J/g °C).

$$q = m \times C \times \Delta T$$
$$m = q/C \times \Delta T$$
$$m = 216 \text{ J} / (0.90 \text{ J/g}^\circ\text{C} \times 20^\circ\text{C})$$
$$m = 12 \text{ g}$$

4. The initial temperature of 150 g of ethanol was 22 °C. What will be the final temperature of the ethanol if 3240 J was needed to raise the temperature of the ethanol? (Specific heat capacity of ethanol is 2.44 J/g °C).

$$q = m \times C \times \Delta T$$
$$\Delta T = q/m \times C$$
$$\Delta T = 3240 \text{ J} / (150 \text{ g} \times 2.44 \text{ J/g}^\circ\text{C})$$
$$\Delta T = 8.85^\circ\text{C}$$
$$T_{\text{final}} = 22^\circ\text{C} + 8.85^\circ\text{C} = 31^\circ\text{C}$$

5. When 435 J of heat is added to 3.4 g of olive oil at 21 °C, the temperature increases to 85 °C. What is the specific heat of the olive oil?

$$q = m \times C \times \Delta T$$
$$C = q/m \times \Delta T$$
$$C = 435 \text{ J} / (3.4 \text{ g} \times 64^\circ\text{C})$$
$$C = 2.0 \text{ J/g}^\circ\text{C}$$

6. How many kilojoules of heat are absorbed when 1.00 L of water is heated from 18 °C to 85 °C? (Hint: You first need to determine the mass of the water, then calculate q in the requested unit.)

Section 3 Reinforcement Using Heat Answers

R Barnett



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Table of Contents Section 3 Reinforcement Using Heat Answers

1. Understanding the eBook Section 3 Reinforcement Using Heat Answers
 - The Rise of Digital Reading Section 3 Reinforcement Using Heat Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 3 Reinforcement Using Heat Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 3 Reinforcement Using Heat Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 3 Reinforcement Using Heat Answers
 - Personalized Recommendations
 - Section 3 Reinforcement Using Heat Answers User Reviews and Ratings
 - Section 3 Reinforcement Using Heat Answers and Bestseller Lists
5. Accessing Section 3 Reinforcement Using Heat Answers Free and Paid eBooks
 - Section 3 Reinforcement Using Heat Answers Public Domain eBooks
 - Section 3 Reinforcement Using Heat Answers eBook Subscription Services
 - Section 3 Reinforcement Using Heat Answers Budget-Friendly Options
6. Navigating Section 3 Reinforcement Using Heat Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Section 3 Reinforcement Using Heat Answers Compatibility with Devices
 - Section 3 Reinforcement Using Heat Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 3 Reinforcement Using Heat Answers
 - Highlighting and Note-Taking Section 3 Reinforcement Using Heat Answers
 - Interactive Elements Section 3 Reinforcement Using Heat Answers
 8. Staying Engaged with Section 3 Reinforcement Using Heat Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 3 Reinforcement Using Heat Answers
 9. Balancing eBooks and Physical Books Section 3 Reinforcement Using Heat Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 3 Reinforcement Using Heat Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Section 3 Reinforcement Using Heat Answers
 - Setting Reading Goals Section 3 Reinforcement Using Heat Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Section 3 Reinforcement Using Heat Answers
 - Fact-Checking eBook Content of Section 3 Reinforcement Using Heat Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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