

Exercise 2: Creating an Activity Series

Table 2: Redox Reactions of Copper, Lead, and Zinc.

Redox Metal	Redox ID	Solution	Immediate Observations	30-Minute Observations
Cu	0	$Pb(NO_3)_2$	No noticeable changes	None
	0	$Zn(NO_3)_2$	No noticeable changes	None
Pb	0	$CuSO_4$	Blue solution, a rust color appeared on the surface	Blue color faded and appeared to be rusty
	0	$ZnSO_4$	None	None
Zn	0	$CuSO_4$	Appeared to be dark, rusty, and fuzzy appearance, blue solution	Appeared to be rusted, was a dark colored almost black, less blue
	0	$Pb(NO_3)_2$	Dark almost black, clear solution	Appeared to be fuzzy, dark, and grey colored, clear solution

Table 3: Potential Redox Reactions and Chemical Equations.

Metal and Aqueous Solution	Reaction Occurred?	Chemical Equation
$Cu + Pb(NO_3)_2$	No	No reaction
$Cu + Zn(NO_3)_2$	No	No reaction
$Pb + CuSO_4$	Yes	$Pb + CuSO_4 \rightarrow Cu + PbSO_4$
$Pb + Zn(NO_3)_2$	No	$Pb + Zn(NO_3)_2 \rightarrow Zn + Pb(NO_3)_2$
$Zn + CuSO_4$	Yes	$Zn + CuSO_4 \rightarrow Cu + ZnSO_4$
$Zn + Pb(NO_3)_2$	Yes	$Zn + Pb(NO_3)_2 \rightarrow Pb + Zn(NO_3)_2$

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Guillaume Favre



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