

A weak base is titrated by a form of the conjugate acid (strong acid). Both are weakly ionizing the equilibrium in the water gives an equilibrium constant K_b . The base equilibrium constant is the reciprocal of the acid equilibrium constant. The water gives off an acid. The water gives off an acid in the process and you have an acid.



Examples:

1. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)

Write the balanced chemical equation for the reaction and label the species.



2. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)

Write the balanced chemical equation for the reaction and label the species.



Answer the following questions. Show all work used to get your answer with units.

1. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____
2. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____
3. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____
4. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____
5. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____
6. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____
7. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____
8. How many mL of a 0.100 M solution of sodium hydroxide must be added to 25.0 mL of a 0.100 M solution of acetic acid to reach the equivalence point? (Ka = 1.8 x 10⁻⁵)
 mL of NaOH = _____

Acid Base Titrations Chem Worksheet 19 5 Answers

K Morrison



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