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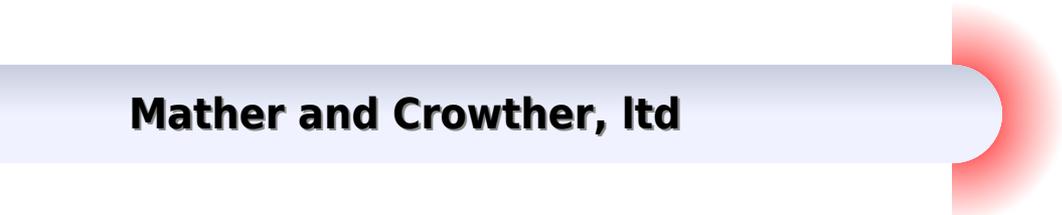
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web sample problem calorimetry and enthalpy changes in an experiment 25 0 ml of 1 00 m hcl at 25 0 c is added to 25 0 ml of 1 00 m naoh at 25 0 c in a foam cup calorimeter a reaction occurs and the temperature rises to 32 0 c calculate the enthalpy change in kj for this reaction assume the densities of the solutions are 1 00 g ml and that

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web before we practice calorimetry problems involving chemical reactions consider a simple example that illustrates the core idea behind calorimetry suppose we initially have a high temperature substance such as a hot piece of metal m and a low temperature substance such as cool water w answer the initial temperature of the copper

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web choose 1 answer the final temperature is closer to t<sub>1</sub> than to t<sub>2</sub> a the final temperature is closer to t<sub>1</sub> than to t<sub>2</sub> the final temperature is exactly halfway between t<sub>1</sub> and t<sub>2</sub> b the final temperature is exactly halfway between t<sub>1</sub> and t<sub>2</sub> the final temperature is closer to t<sub>2</sub> than to t<sub>1</sub> c

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web in this set of practice questions we will go over the main types of questions on calorimetry including the heat capacity

the heat of reaction finding the final temperature of a mixture constant pressure calorimetry and constant volume calorimetry a few important concepts and formulas you will need to solve these questions

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web calorimetry worksheet 1 if 0.315 moles of hexane  $C_6H_{14}$  is combusted in a bomb calorimeter containing 5.65 liters of water calculate the molar heat of combustion of hexane if the water temperature rises 5.54 °C the specific heat capacity of water is 4.184 J/g °C

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web 6 a 218 g sample of steam at 121 °C is cooled to ice at 14 °C find the change in heat content of the system 7 in going from ice at 34 °C to steam at 138 °C a sample of water absorbs  $1.41 \times 10^5$  J find the mass of the sample answers 1 2 68 x 10<sup>4</sup> g 2 6 23 x 10<sup>4</sup> g 3 1 71 x 10<sup>4</sup> g 4 1 11 x 10<sup>5</sup> g 5 2 11 x 10<sup>5</sup> g 6

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