

Calculate the final pH of a buffer containing 0.1 M sodium acetate and 0.2 M acetic acid ( $pK_a = 4.76$ ) following the addition of 0.05 M NaOH.

$$pH = pK_a + \log \frac{\text{Base}}{\text{Acid}}$$

$$pH = 4.76 + \log \frac{0.1 + 0.05}{0.2 - 0.05}$$

$$pH = 4.76 + \log \frac{0.15}{0.15}$$

$$pH = 4.76$$

# Ph Properties Of Buffer Solutions Lab Calculations

**Yan Bai**



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In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Ph Properties Of Buffer Solutions Lab Calculations**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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