

Atomic bonding

Ionic: Anions gain electrons, cations lose electrons

Covalent: electrons shared between two atoms

Metallic: electrons shared between many atoms

Van der Waals: temporary or permanent dipole

Interacts with other objects

Crystal structures

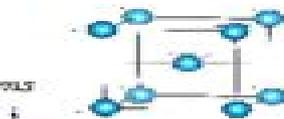
BODY CENTERED CUBIC

$$4R = \sqrt{a^2 + a^2 + a^2}$$

$$4R = a\sqrt{3}$$

1 center atom, 8 corner atoms

2 atoms per unit cell



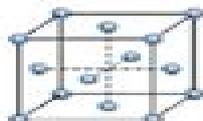
FACE CENTERED CUBIC

$$4R = \sqrt{a^2 + a^2}$$

$$4R = a\sqrt{2}$$

6 face atoms, 8 corner atoms

4 atoms per unit cell



Volume of a cubic = a^3

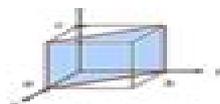
$$A = 10^{-10} \text{ m}$$

$$\text{Avogadro No} = 6.022 \times 10^{23} \left(\frac{\text{atoms}}{\text{mol}} \right)$$

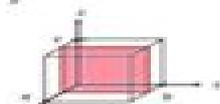
$$\text{Density} = \frac{\text{mass}}{\text{volume}} = \frac{\# \text{ of atoms} \times \text{atomic mass}}{\text{volume of unit cell} \times \text{Avogadro}}$$

Miller Indices

crystal	h	k	l
1. Intercepts	1	1	1
2. Reciprocals	1/1	1/1	1/1
3. Reduction	1	1	1
4. Miller indices	(111)		



crystal	h	k	l
1. Intercepts	1/2	1	1
2. Reciprocals	2/1/2	1/1	1/1
3. Reduction	2	1	2
4. Miller indices	(212)		



$$\text{Planar Density} = \frac{\text{atoms per face}}{\text{area of face}}$$

$$\text{Packing Factor} = \frac{\text{volume of atom} \times \text{number of atoms}}{\text{volume of the unit cell}}$$

Point Defects

$$n_v = n \cdot e^{-\frac{Q_v}{RT}}$$

$$n_v = \# \frac{\text{vacancies}}{\text{cm}^3}, n = \# \frac{\text{atoms}}{\text{cm}^3}$$

Q_v = energy for 1 mol of vacancies, T = Temp

$$R = \text{Gas constant} = 1.987 \frac{\text{cal}}{\text{mol} \cdot \text{K}} = 8.314 \frac{\text{joules}}{\text{mol} \cdot \text{K}}$$

Diffusion

$$D = D_0 \cdot e^{-\frac{Q}{RT}}$$

D_0 = constant based on material

Q_v = energy for 1 mol of vacancies, T = Temp

$$R = \text{Gas constant} = 1.987 \frac{\text{cal}}{\text{mol} \cdot \text{K}} = 8.314 \frac{\text{joules}}{\text{mol} \cdot \text{K}}$$

$$J = -D \cdot \frac{dc}{dx}, x = \text{thickness}, J = \text{Flux} \text{ - Fick's 1st law}$$

c = Compositions in atomic percent

Mechanical Properties

$$\text{Engineering stress (S)} = \frac{F}{A_0}$$

$$\text{Engineering Strain (e)} = \frac{\Delta l}{l}$$

$$\text{Young Modulus (E)} = \frac{S}{e}$$

F = Force, A_0 = Area

Δl = Change in length, l = Original length

$$K_f \text{ or } K_{fc} = f \sqrt{\pi a}$$

σ = Exposed stress, a = flaw size

f = geometry factor for the specimen and flaw

K_{fc} = fracture toughness

K_{Ic} = plain strain fracture toughness

$$\pm \sigma = 5.09 \cdot \frac{FL}{d^3}$$

F = Applied Load, L = Length of shaft

d = diameter of shaft

$$1 - K(V) = \exp \left[- \left(\frac{\sigma}{\sigma_0} \right)^m \right]$$

m = Weibull Modulus, σ_0 = Flexural Strength

F = Failure, σ_0 = Initial Flexural Strength

$$\text{Poissons Ratio (v)} = \frac{-\epsilon_{lateral}}{\epsilon_{longitudinal}}$$

$$\text{Modulus of resilience (E}_r) = \frac{1}{2} (\text{yield strength}) (\text{strain@yielding})$$

$$\text{Ductility \%Elongation} = \frac{l_f - l_0}{l_0} \cdot 100\%$$

$$\text{Ductility \%Reduction in area} = \frac{A_0 - A_f}{A_0} \cdot 100\%$$

$$\text{True Stress} = \frac{F}{A_f}$$

$$\text{True Strain} = \ln \left(\frac{l}{l_0} \right)$$

$$\sigma_t = K \epsilon_t^n$$

n = strain coefficient, K = Constant

ϵ_t = True strain, σ_t = True Stress

$$\sigma_m = 2\sigma_0 \left(\frac{a}{r_0} \right)^{\frac{1}{2}}$$

σ_m = Stress @ Crack tip, σ_0 = Applied stress

a = Crack Length, r_0 = radius of curvature

$$\sigma_c = \left(\frac{2E\gamma_s}{\alpha\pi} \right)^{\frac{1}{2}}$$

σ_c = Critical Stress, E = elastic modulus

γ_s = specific surface energy

α = one half length of internal crack

Solidification and Nucleation

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28 full papers and 17 short papers included in this book were carefully reviewed and selected from 109 submissions They were organized in topical sections as follows Full Conference Papers Work in Progress Interactive Entertainment Showcase and Student Competition *Innovations in E-learning, Instruction Technology, Assessment and Engineering Education* Magued Iskander,2007-09-04 This book includes a set of rigorously reviewed world class manuscripts addressing and detailing state of the art research projects in the areas of Engineering Education Instructional Technology Assessment and E learning The book presents selected papers form the conference proceedings of the International Conference on Engineering Education Instructional Technology Assessment and E learning EIAE 2006 All aspects of the conference were managed on line **National Educators' Workshop: Update 2002 - Standard Experiments in Engineering, Materials Science, and Technology** Edwin J. Prior,2003 Earth Science Demystified Linda D. Williams,2004-07-14 Say goodbye to dry presentations grueling formulas and abstract theories that would put Einstein to sleep now there s an easier way to master the disciplines you really need to know McGraw Hill s Demystified Series teaches complex subjects in a unique easy to absorb manner and is perfect for users without formal training or unlimited time They re also the most time efficient interestingly written brush ups you can find Organized as self teaching guides they come complete with key points background information questions at the end of each chapter and even final exams You ll be able to learn more in less time evaluate your areas of strength and weakness and reinforce your knowledge and confidence Earth Science has never been easier to understand Coverage includes rocks and minerals strata fossils volcanos earthquakes glaciers wind and erosion oceans type of rock atmosphere carbon and calcium the hydrologic cycle and more Matriculation model answers: French. Being the papers from June 1888 to Jan. 1894 (to Jan. 1897, June 1893 to Jan. 1899, to Jan. 1900, to Jan. 1902, June 1897 to Sept. 1902, Jan. 1899 to Sept. 1904, to Sept. 1905, to Sept. 1906, Jan. 1901 to Jan. 1907, to Sept. 1907, to Sept. 1908, Jan. 1905 to June 1909, to Jan. 1910, to Sept. 1910, to June 1911, June 1911 to June 1915), with answers by tutors of University correspondence college London univ, exam. papers,1894 How We Live: Test Packet Clp71512,2006-09-06 The Science Teacher ,2007 **Chemical Discipline-specific Review for the FE/EIT Exam** Stephanie T. Lopina,1997 The best preparation for discipline specific FE exams 60 practice problems with full solutions Two complete simulated 4 hour discipline specific exam Covers all the topics for that particular discipline Provides the in depth review you need Topics covered Chemical Reaction Engineering Chemical Thermodynamics Computers Numerical Methods Heat Transfer Mass Transfer Material Energy Balances Pollution Prevention Process Control Process Design Economics Evaluation Process Equipment Design Process Safety Transport Phenomena *Mechanical Discipline-specific Review for the FE/EIT Exam* Michel A. Saad,1997 The FE exam the first in the two part engineering licensing process is taken typically by upper level students or recent graduates in April or October This eight hour exam is closed book except for a handout provided in the examination room The exam is divided into morning and afternoon sessions The morning exam with 120 multiple choice

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