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550 j  $\Delta h$  q water 550 j

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calorimeter with a total heat capacity of 15.48 kJ/C the temperature of the calorimeter and its contents changes from 19.77°C to 33.90°C calculate  $q$  and  $h$  for the following reaction as written  $2\text{NH}_3(\text{g}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g}) + 2\text{CH}_4(\text{g})$

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web choose 1 answer the final temperature is closer to  $t_1$  than to  $t_2$  a the final temperature is closer to  $t_1$  than to  $t_2$  the final temperature is exactly halfway between  $t_1$  and  $t_2$  b the final temperature is exactly halfway between  $t_1$  and  $t_2$  the final temperature is closer to  $t_2$  than to  $t_1$  c

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web 6 a 218 g sample of steam at 121°C is cooled to ice at 14°C find the change in heat content of the system 7 in going from ice at 34°C to steam at 138°C a sample of water absorbs  $1.41 \times 10^5$  J find the mass of the sample answers 1 2 68  $\times 10^4$  J 2 6 23  $\times 10^4$  J 3 1 71  $\times 10^4$  J 4 1 11  $\times 10^5$  J 5 2 11  $\times 10^5$  J 6

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