

# KINETICS Practice Problems and Solutions

Name: \_\_\_\_\_  
 Period: \_\_\_\_\_ Date: \_\_\_\_\_

AP Chemistry  
 Dr. Mandes

The following questions represent potential types of quiz questions. Please answer each question completely and thoroughly. The solutions will be posted on-line on Monday.

5. Please do #18 in chapter 12 of your text.
- This increases the  $[H_2]$ , which will increase the rate, but has no effect on  $k$
  - Due to Arrhenius equation, changing temperature changes the value of  $k$ .
  - Catalyst function by lowering the activation energy, so due to Arrhenius equation, changing the activation energy changes the value of  $k$ .

6. Consider the reaction:  $P_4 + 6 H_2 \rightarrow 4 PH_3$ . A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[P_4]$ , mol/L	$[H_2]$ , mol/L	Initial Rate, mol/L · s
0.0110	0.0075	$3.20 \times 10^{-4}$
0.0110	0.0150	$6.40 \times 10^{-4}$
0.0220	0.0150	$6.39 \times 10^{-4}$

- What is the order with respect to:  $P_4$  0  
 $H_2$  1
- Write the rate law for this reaction.  $rate = k[H_2]$
- Determine the value and units of the rate constant,  $k$ . **plug and chug using the rate law & data from exp't 1 and solving for  $k$ , we get  $k = 0.0427 s^{-1}$**

7. Consider the reaction:  $SO_2 + O_3 \rightarrow SO_3 + O_2$ . A rate study of this reaction was conducted at 298 K. The data that were obtained are shown in the table.

$[SO_2]$ , mol/L	$[O_3]$ , mol/L	Initial Rate, mol/L · s
0.25	0.40	0.118
0.25	0.20	0.118
0.75	0.20	1.062

- What is the order with respect to:  $SO_2$  2  
 $O_3$  0
- Write the rate law for this reaction.  $rate = k[SO_2]^2[O_3]^0$
- Determine the value and units of the rate constant,  $k$ . **plug and chug using the rate law & data from exp't 1 and solving for  $k$ , we get  $k = 2.36 molL^{-2} s^{-1}$**

8. Consider the following mechanism.
- $$A_2 + B_2 \rightarrow R + C \quad (\text{slow})$$
- $$A_2 + R \rightarrow C \quad (\text{fast})$$

- Write the overall balanced chemical equation.  $2 A_2 + B_2 \rightarrow 2 C$
- Identify any intermediates within the mechanism. **R**
- What is the order with respect to each reactant?  **$A_2$  1<sup>st</sup>;  $B_2$  1<sup>st</sup>**

# Solutions Chemical Kinetics

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**Chemical Kinetics** Kenneth Antonio Connors, 1990 *Chemical Kinetics The Study of Reaction Rates in Solution* Kenneth A Connors This chemical kinetics book blends physical theory phenomenology and empiricism to provide a guide to the experimental practice and interpretation of reaction kinetics in solution It is suitable for courses in chemical kinetics at the graduate and advanced undergraduate levels This book will appeal to students in physical organic chemistry physical inorganic chemistry biophysical chemistry biochemistry pharmaceutical chemistry and water chemistry all fields concerned with the rates of chemical reactions in the solution phase *Solutions Manual: Sm Chemical Kinetics and React Dyn* Paul L. Houston, 2001

**Reaction Kinetics** Keith J. Laidler, 2013-10-22 *Reaction Kinetics Volume II Reactions in Solution* deals with the kinetics of reactions in solution and discusses the basic principles and theories of kinetics including a brief description of homogeneous gas reactions This book is divided into two chapters The first chapter focuses on the general principles of reactions in solution that includes reactions between ions and involving dipoles influence of pressure on rates in solution substituent effects and homogeneous catalysis in solution Chapter 2 primarily deals with general features of reactions in solution emphasizing the relationship between the results of a kinetic investigation and actual reaction mechanism This volume is intended for undergraduate students of chemistry who have not previously studied chemical kinetics This book is also useful to more advanced students in other fields such as biology and physics who wish to have a general knowledge of the subject

**International Journal of Chemical Kinetics**, 1993 Topics covered in this publication include quantitative relationships between molecular structure and chemical activity organic inorganic chemistry biochemical kinetics and reaction mechanisms Surface kinetics are also explored

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**Classification for Works on Pure and Applied Science in the Science Library, the Science Museum** Science Museum (Great Britain). Library, 1921 *Scientific and Technical Aerospace Reports*, 1966 Monthly Catalog of United States Government Publications, **The Kinetics of Reactions in Solution** Emyr Alun Moelwyn-Hughes, 1947 Problems and Solutions to Chemical Kinetics and Reaction Dynamics Paul L. Houston, 1995 Science Progress, 1916

**British Books in Print**, 1924 **Monthly Catalog of United States Government Publications, Cumulative Index** United States. Superintendent of Documents, 1976 **A System of Physical Chemistry: Kinetic theory** William Cudmore

McCullagh Lewis, 1918

**General Catalogue** Carnegie Institute of Technology, 1915

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