

Basic Stoichiometry Post-Lab Homework Exercises

1. Load the "Reactants, Products, and Leftovers" simulation and work through each of the levels of the Game! At <http://phet.colorado.edu/> or googling "phet." You may have to download or update Java on your computer.

Show your work for all calculations. Remember: Labels, units, sig figs and cancellation!

2. For the reaction $P_4 + 6Cl_2 \rightarrow 4PCl_3$, determine how many moles of chlorine Cl_2 would be needed to react with 4.0 moles of phosphorus P_4 to entirely use up all the phosphorus. 4) _____

3. If 3.00 moles of P_4 reacted with 22.0 moles Cl_2 according to the above reaction, determine:

- a) How many moles PCl_3 are produced a) _____
- b) How many moles of P_4 are left in excess after the reaction (if any) b) _____
- c) How many moles of Cl_2 are left in excess after the reaction (if any) c) _____

In reality, reactants don't have to react in perfect whole-numbers of moles. In a two-reactant synthesis reaction, usually **one reactant gets entirely used up (and determines how much product is made)**, even if that means using fractions of a mole of reactant. For instance, when solid, metallic aluminum Al and red, liquid bromine Br_2 are brought together, they make a white solid according to the reaction $2Al + 3Br_2 \rightarrow 2AlBr_3$. If 5.0 moles of aluminum Al was reacted with 10 moles bromine Br_2 , all five moles of aluminum would react, with only 7.5 moles bromine. (2:3 mole ratio) This would produce only 5.0 moles of $AlBr_3$, leaving 2.5 moles of excess Br_2 behind.

5. Now assume 3.0 moles Al and 4.0 moles Br_2 react.

- a) How much (in moles) $AlBr_3$ gets produced? a) _____
- b) Which chemical must be the *excess reactant*? b) _____
- c) Which chemical is the *limiting reactant*? c) _____
- d) If all the limiting reactant gets used up, how much of the excess reactant is left? d) _____

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Answer Key

PT Brinkman



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Table of Contents Basic Stoichiometry Post Lab Homework Exercises Answer Key

1. Understanding the eBook Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - The Rise of Digital Reading Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - Personalized Recommendations
 - Basic Stoichiometry Post Lab Homework Exercises Answer Key User Reviews and Ratings
 - Basic Stoichiometry Post Lab Homework Exercises Answer Key and Bestseller Lists
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 - Basic Stoichiometry Post Lab Homework Exercises Answer Key Compatibility with Devices
 - Basic Stoichiometry Post Lab Homework Exercises Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - Highlighting and Note-Taking Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - Interactive Elements Basic Stoichiometry Post Lab Homework Exercises Answer Key
8. Staying Engaged with Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Basic Stoichiometry Post Lab Homework Exercises Answer Key
9. Balancing eBooks and Physical Books Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Basic Stoichiometry Post Lab Homework Exercises Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Setting Reading Goals Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Basic Stoichiometry Post Lab Homework Exercises Answer Key
- Fact-Checking eBook Content of Basic Stoichiometry Post Lab Homework Exercises Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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