

## Chapter 5 Chapter Review Question Answers

- Describe the scientists' contributions:
  - Cannizzaro—developed a standard method for measuring atomic masses, which allowed chemists to search for periodic trends among elements.
  - Mendeleev—organized elements according to increasing atomic mass and noticed that similar properties appeared periodically.
  - Moseley—discovered that atomic number, not atomic mass, should be the basis of organizing elements on the periodic table.
- The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- Groups of elements exhibit similar chemical and physical properties and behaviors.
- Electron configurations:
  - Generally, the configurations of the outermost electrons shells of elements within the same group are the same. There are a number of exceptions to this rule, however, among the transition elements.
  - Their outer shells are completely filled.
- The length of a period is determined by the total number of electrons that can fill the outer sublevels of the elements of that period.
- An element's period corresponds to its highest occupied main energy level. (s & p)
- Blocks & Groups
  - The type of sublevel being filled in successive elements of that block.
  - s-block: Groups 1 & 2; p-block: Groups 13–18; d-block: Groups 3–12; f-block: No group #s, as they are between groups 3 & 4, but period 6 elements are in the Lanthanide series, and period 7 elements are in the Actinide series.
- Alkali Metals
  - The Group 1 elements, except Hydrogen.
  - Alkali metals are extremely reactive, they react violently with water, they have low densities, they are soft enough to be cut with a knife, and freshly cut alkali metals are very shiny silver colored metals.
- Alkaline Metals
  - The Group 2 elements.
  - They are slightly less reactive than Group 1 elements; they are harder, duller, denser than Group 1 elements. They also have slightly higher melting points than the Group 1 elements.
- Group Configurations:
  - Transition Metals (d-block)
    - Group 3:  $ns^2 (n-1)d^1$
    - Group 4:  $ns^2 (n-1)d^2$
    - Group 5:  $ns^2 (n-1)d^3$
    - Group 6:  $ns^2 (n-1)d^4$  {actually is  $ns^1 (n-1)d^5$  for Cr & Mo}
    - Group 7:  $ns^2 (n-1)d^5$
    - Group 8:  $ns^2 (n-1)d^6$
    - Group 9:  $ns^2 (n-1)d^7$
    - Group 10:  $ns^2 (n-1)d^8$
    - Group 11:  $ns^2 (n-1)d^9$  {actually is  $ns^1 (n-1)d^{10}$  for all group 11 elements}
    - Group 12:  $ns^2 (n-1)d^{10}$
  - The group number is the sum of the s & d electrons.
- The transition elements.
- p-block
  - The p-block consists of nonmetals at the right, metalloids in the middle along the staircase, and metals at the left.
  - The p-block metals are generally harder and denser than the s-block metals, but softer and less dense than the d-block metals.
- Halogens
  - The Group 17 elements.

# Modern Chemistry Chapter 5 Review Answer Key

**Arturo Cuomo**



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### **Table of Contents Modern Chemistry Chapter 5 Review Answer Key**

1. Understanding the eBook Modern Chemistry Chapter 5 Review Answer Key
  - The Rise of Digital Reading Modern Chemistry Chapter 5 Review Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modern Chemistry Chapter 5 Review Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modern Chemistry Chapter 5 Review Answer Key
  - Personalized Recommendations
  - Modern Chemistry Chapter 5 Review Answer Key User Reviews and Ratings
  - Modern Chemistry Chapter 5 Review Answer Key and Bestseller Lists
5. Accessing Modern Chemistry Chapter 5 Review Answer Key Free and Paid eBooks

- Modern Chemistry Chapter 5 Review Answer Key Public Domain eBooks
  - Modern Chemistry Chapter 5 Review Answer Key eBook Subscription Services
  - Modern Chemistry Chapter 5 Review Answer Key Budget-Friendly Options
6. Navigating Modern Chemistry Chapter 5 Review Answer Key eBook Formats
- ePub, PDF, MOBI, and More
  - Modern Chemistry Chapter 5 Review Answer Key Compatibility with Devices
  - Modern Chemistry Chapter 5 Review Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Modern Chemistry Chapter 5 Review Answer Key
  - Highlighting and Note-Taking Modern Chemistry Chapter 5 Review Answer Key
  - Interactive Elements Modern Chemistry Chapter 5 Review Answer Key
8. Staying Engaged with Modern Chemistry Chapter 5 Review Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modern Chemistry Chapter 5 Review Answer Key
9. Balancing eBooks and Physical Books Modern Chemistry Chapter 5 Review Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Modern Chemistry Chapter 5 Review Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Modern Chemistry Chapter 5 Review Answer Key
- Setting Reading Goals Modern Chemistry Chapter 5 Review Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modern Chemistry Chapter 5 Review Answer Key
- Fact-Checking eBook Content of Modern Chemistry Chapter 5 Review Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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