

First 100 High Frequency Words

in frequency order reading down the columns

the	that	not	look	put
and	with	then	don't	could
a	all	were	come	house
to	we	go	will	old
said	can	little	into	too
in	are	as	back	by
he	up	no	from	day
I	had	mum	children	made
of	my	one	him	time
it	her	them	Mr	I'm
was	what	do	get	if
you	there	me	just	help
they	out	down	now	Mrs
on	this	dad	came	called
she	have	big	oh	here
is	went	when	about	off
for	be	it's	got	asked
at	like	see	their	saw
his	some	looked	people	make
but	so	very	your	an

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Joseph Green



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web elements in the second row of the periodic table place their electrons in the $2n$ shell as well as the $1n$ shell for instance lithium li has three electrons two fill the $1s$ orbital and the third is placed in the $2s$ orbital giving an electron configuration of $1s^2 2s^1$

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web electron configurations and trends answer these discussion questions given the following part of an electron

configuration $3s^2$ explain which part refers to the number of electrons the energy level and the sublevel what is the maximum number of electrons that can occupy a single orbital state and explain Hund's rule

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web one more point needs to be emphasized about the relationship between electron configuration and the periodic table the atoms of elements in the same vertical column of the table have similar electron configurations for example consider the alkaline earth elements group IIA

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web the electron removed during the ionization of beryllium $He\ 2s^2$ is an s electron whereas the electron removed during the ionization of boron $He\ 2s^2\ 2p^1$ is a p electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are $He\ 2$ helium $1s^2$ $Ne\ 10$ neon $1s^2\ 2s^2\ 2p^6$

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web q1 what is the electron configuration of the element potassium answer the atomic number of potassium is 19 a neutral potassium element has 19 electrons its electron configuration is $1s^2\ 2s^2\ 2p^6\ 3s^2\ 3p^6\ 4s^1$ or simply put as $4s^1$ q2 what is the total number of electrons that can be filled in s p and d subshells

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