

Answer KEY

Unit 3 Outline/ Study Guide - Electron Configuration and Periodic Trends

Essential Skills/ State Standards:

1. Students know how to relate the position of an element in the periodic table to its quantum electron configuration and to its reactivity with other elements in the table.
2. Students know how to use the periodic table to identify metals, semimetals, nonmetals, halogens, alkali metals, alkaline earth metals, and transition metals.
3. Students know how to use the periodic table to identify the following trends:
 - ionization energy, electronegativity, and the relative sizes of ions and atoms
 (this includes knowing the definitions of ionization energy, electronegativity, cation, and anion.)

Study Guide - (use your notes and worksheets for a more complete review)

1. Identify the following things on the periodic table:

-metals	-alkali metals
-metalloids (semi-metals)	-alkaline earth metals
-transition metals	-halogens
-nonmetals	-noble gases
-Periods	-Groups/families

Identify groups with the following numbers of valence electrons:
1, 2, 3, 4, 5, 6, 7, 8

Key:
 Metals (red)
 Metalloids (purple)
 Nonmetals (blue)

Group 1: Alkali metals
 Group 2: Alkaline earth metals
 Groups 3-10: Transition metals
 Group 11: Coinage metals
 Group 12: Zinc group
 Group 13: Boron group
 Group 14: Carbon group
 Group 15: Nitrogen group
 Group 16: Chalcogens
 Group 17: Halogens
 Group 18: Noble gases

2. Why do the elements in the groups (columns) behave so similar to each other even if they have great differences in their number of protons? **same number of valence electrons and similar chemical properties.**
3. What is the connection between the following terms: **octet rule, valence electrons, and ions?** In order to achieve the **octet rule**, atoms **gain or lose their valence electrons**. They become charged **ions** when this occurs.
4. a. List the elements that are exceptions to the octet rule b) Why don't these atoms achieve the octet rule?
 - a. Small atoms are happy with 2 electrons (ex: H, He, Li, Be, B)
 - b. because the 1s energy level only holds 2 electrons

Atoms Periodic Table Review Answer Key Lbrsfs

George Graybill



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Atoms, Molecules & Elements: What Are Atoms? Gr. 5-8 George Graybill, 2015-10-01 This is the chapter slice What Are Atoms from the full lesson plan Atoms Molecules Elements Young scientists will be thrilled to explore the invisible world of atoms molecules and elements Our resource provides ready to use information and activities for remedial students using simplified language and vocabulary Students will label each part of the atom learn what compounds are and explore the patterns in the periodic table of elements to find calcium Ca chlorine Cl and helium He through hands on activities These and more science concepts are presented in a way that makes them more accessible to students and easier to understand Written to grade and using simplified language and vocabulary and comprised of reading passages student activities crossword word search comprehension quiz and color mini posters our resource can be used effectively for test prep and your whole class All of our content is aligned to your State Standards and are written to Bloom s Taxonomy and STEM initiatives

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