

Detection of Ions in Solutions Using Acid/Base Chemistry: A Quality Control Test

Objective: This lab focuses on the detection of ions using titration as an analysis tool. You will standardize NaOH and HCl solutions so that you know the exact concentration and then prepare samples of common household items in order to determine the amount of calcium in Tang[®], Mg(OH)₂ in Milk of Magnesia, etc. You will learn to prepare samples of a specified concentration, learn about acids and bases through the use of titrations and learn how to detect endpoints using different indicators. You will become adept at measuring pH with both a pH meter and indicator paper. You will then conduct quality control testers and determine if the label on a bottle of over-the-counter product actually contains the percentage of compound that it advertises.

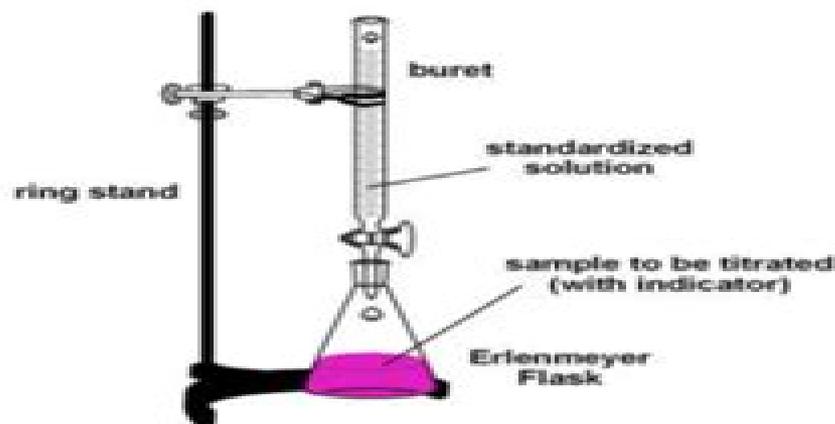
What is a Titration?

A titration is an analytical procedure used to determine the concentration of a sample by reacting it with a standard solution. One type of titration uses a neutralization reaction, in which an acid and a base react to produce a salt and water.



In equation 1, the acid is HCl (called hydrochloric acid) and the base is NaOH (called sodium hydroxide). When the acid and base react, they form NaCl (sodium chloride), which is also known as table salt. The titration proceeds until the equivalence point is reached, where the number of moles of acid is equal to the number of moles of base. This point is usually marked by observing a color change in an added indicator.

In a titration, the standard solution goes in a buret, which is a piece of glassware used to measure the volume of solvent to approximately 0.1 mL of accuracy. The solution that you are titrating goes in an Erlenmeyer flask, which should be large enough to accommodate both your sample and the standard solution you are adding.



Titration Lab Answers

**Jeffrey Stanton, Robert P. Wells, Sandra
Rochowansky**

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