

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then use data in the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount of NO_2 Excess Reactant
1 molecule	1 molecule	1 molecule	none	none
2 molecules	2 molecules	2 molecules	NO	1 molecule O_2
2 molecules	3 molecules	4 molecules	O_2	2 molecules NO
1.00 mol	2.00 mol	2.00 mol	NO	1.00 mol
2.00 mol	4.00 mol	4.00 mol	NO	2.00 mol O_2
3.00 mol	6.00 mol	6.00 mol	NO	3.00 mol O_2
4.00 mol	8.00 mol	8.00 mol	NO	4.00 mol O_2
5.00 mol	10.00 mol	10.00 mol	NO	5.00 mol O_2
6.00 mol	12.00 mol	12.00 mol	NO	6.00 mol O_2
7.00 mol	14.00 mol	14.00 mol	NO	7.00 mol O_2
8.00 mol	16.00 mol	16.00 mol	NO	8.00 mol O_2
9.00 mol	18.00 mol	18.00 mol	NO	9.00 mol O_2
10.00 mol	20.00 mol	20.00 mol	NO	10.00 mol O_2
10.00 g	30.01 g	20.00 g	O_2	10.01 g NO
10.00 g	60.02 g	40.01 g	O_2	20.02 g NO
10.00 g	90.03 g	60.02 g	O_2	30.03 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

32.00 g O_2 = 1 mol O_2 /32.00 g O_2 = 0.3125 mol O_2

30.00 g NO = 1 mol NO /30.01 g NO = 0.9998 mol NO

actual mole ratio = 0.9998 mol NO /0.3125 mol O_2 = 3.197 mol NO /1.000 mol O_2

Because the actual mole ratio of NO / O_2 is larger than the balanced equation mole ratio of NO / O_2 , there is an excess of NO / O_2 is the limiting reactant.

Mass of NO used = 0.3125 mol O_2 = 2 mol NO /1 mol O_2 = 0.6250 mol NO

0.6250 mol NO = 30.01 g NO /1 mol NO = 18.76 g NO

Mass of NO_2 produced = 0.6250 mol NO_2 = 46.01 g NO_2 /1 mol NO_2 = 28.76 g NO_2

Excess NO = 20.00 g NO - 18.76 g NO = 1.24 g NO

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CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium nitride, a substance used in the manufacturing of some semiconductors for lasers.



In one experiment, the reaction yielded 7.42 g of the oxide from a 1.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and GaN_3 are 69.72 g/mol and 127.63 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield volume = 7.42 g GaN_3
2. mass of reactant = 1.00 g Ga, 100% Ga
3. number of moles of reactant = 1.00 g Ga \times $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$ = 7.00 g Ga = 1 mol Ga/69.72 g Ga
4. number of moles of product = 0.00 mol Ga \times $\frac{2 \text{ mol } GaN_3}{1 \text{ mol Ga}}$ = $\frac{0.100 \text{ mol Ga} \times 127.63 \text{ g Ga}}{2 \text{ mol } GaN_3}$ = 6.38 mol Ga
5. theoretical yield = 0.0000 mol GaN_3 \times $\frac{127.63 \text{ g } GaN_3}{1 \text{ mol } GaN_3}$ = correct
6. percent yield = $\frac{7.42 \text{ g } GaN_3}{127.63 \text{ g } GaN_3} \times 100$ = $\frac{7.42 \text{ g } GaN_3}{6.38 \text{ g } GaN_3} \times 100$ = 166

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