

Solutions for Gravimetric Analysis Exercises

1. The terms in a reaction quotient are actually dimensionless ratios of actual concentrations (or pressures) divided by standard concentrations (or pressures). The standard state for solutes is a 1 M solution and for gases it is a pressure of 1 bar (~ 1 atm), so these are the units used. For liquids and solids the standard state is the pure substance, so these ratios are unity. The solvent in a dilute solution is approximated as a pure substance.
2. The equilibrium constant for a reaction is dimensionless because it equals the reaction quotient at equilibrium, and, as discussed in Problem 1, this ratio involves only dimensionless terms.

3. a) Adding the two equations together gives



- b) At equilibrium $K_3 = [\text{AgCl}(\text{aq})]$, so $[\text{AgCl}(\text{aq})] = 3.6 \times 10^{-7}$ M.

4. Determine the solubility of $\text{La}(\text{IO}_3)_3$ using K_{sp} for $\text{La}(\text{IO}_3)_3$ and a table of initial and equilibrium concentration terms.



where x equals the increase in the $[\text{La}^{3+}]$ as a result of the dissolving process.

$$\text{at equil, } K_{sp} = 1.0 \times 10^{-11} = [\text{La}^{3+}][\text{IO}_3^-]^3 = (x)(0.050 + 3x)^3 = (x)(0.050)^3$$

\downarrow
 assumes $3x \ll 0.050$

$$[\text{La}^{3+}] = \frac{1.0 \times 10^{-11}}{(0.050)^3} = 8.0 \times 10^{-9} \text{ M} \quad (\text{assumption valid})$$

Solutions For Gravimetric Analysis Exercises

Alexander Humboldt Sexton



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Elementary Chemistry Harry Mann Gordin, 1913 **Exercises in Elementary Quantitative Chemical Analysis for Studentsof Agriculture** Azariah Thomas Lincoln, James Henri Walton, 1907 *Exercises in Elementary Quantitative Chemical Analysis for Students of Agriculture* Azariah Thomas Lincoln, James Henri Walton, 1907 KVPY - SA : Chemistry for Class 11th by Career Point Kota Career Point Kota, 2020-07-26 Career Point Kota feel great pleasure to present before you this KVPY SA book Detailed Topic Wise theory supported with example Previous Year Questions Complete Solution This book is designed for the aspirants of KVPY Stream SA As there is no prescribed syllabus for KVPY hence this books is designed considering the topics from where questions have been asked in previous years The book is scientifically structured to prepare aspirants of KVPY Each chapter has detailed topic wise Theory supported with examples to understand the application of concepts followed by Exercise 1 covering the different patterns of questions to give sufficient practice to the students After this Exercise 2 is given covering previous years questions to give exposure to type of questions asked Complete solutions of exercise sheets are also provided in the book itself These solutions are not just sketch rather have been written in such a manner that the students will be able to understand the application of concept and can answer some other related questions too We firmly believe that the book in this form will definitely help a genuine hardworking student We have tried our best to keep errors out of this book Comment and criticism from readers will be highly appreciated and incorporated in the subsequent edition We wish to utilize the opportunity to place on record our special thanks to all team members of Content Development for their efforts to make this wonderful book *A Text-book of Quantitative Chemical Analysis* Alexander Charles Cumming, Sydney Alexander Kay, 1919 Chemistry Rhobert Lewis, Wynne Evans, 2018-03-31 The fifth edition of this engaging and established textbook provides students with a complete course in chemical literacy and assumes minimal prior experience of science and maths Written in an accessible and succinct style this book offers comprehensive coverage of all the core topics in organic inorganic and physical chemistry Topics covered include bonding moles solutions and solubility energy changes equilibrium organic compounds and spectroscopy Each unit contains in text exercises and revision questions to consolidate learning at every step and is richly illustrated with diagrams and images to aid understanding This popular text is an essential resource for students who are looking for an accessible introductory textbook It is also ideal for non specialists on courses such as general science engineering environmental health or life sciences New to this Edition A foreword by Professor Sir John Meurig Thomas FRS former Director of the Royal Institution Three additional units on Gibbs Energy Changes Organic Mechanisms and Fire and Flame Accompanying online resources for this title can be found at bloomsburyonlineresources.com/chemistry/5e These resources are designed to support teaching and learning when using this textbook and are available at no extra cost **Analytical Chemistry** Frederick Pearson Treadwell, 1928 *Practical Chemistry by Micro-methods* Egerton Charles Grey, 1925 A Text-book of Quantitative

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