

In Table 4 below,  $[\text{Fe}^{3+}]$  after mixing is obtained by using  $M_1V_1 = M_2V_2$ . This data is to be transferred to Table 5 below as the initial  $[\text{Fe}^{3+}]$  (column 2).

Similarly in Table 4 of Data sheet  $[\text{SCN}^-]$  after mixing obtained using  $M_1V_1 = M_2V_2$ . This data is then to be transferred to Table 5 below as the initial  $[\text{SCN}^-]$  (column 3).

Table 4. Initial Concentration Data

Test Tube	$[\text{Fe}^{3+}]$ Before Mixing	$[\text{Fe}^{3+}]$ After Mixing <sup>1</sup> (Initial $[\text{Fe}^{3+}]$ )	$[\text{SCN}^-]$ Before Mixing	$[\text{SCN}^-]$ After Mixing <sup>1</sup> (Initial $[\text{SCN}^-]$ )
F	0.01	0.0025M	0.00113	$2.825 \times 10^{-4}$
G	0.01	0.0025M	0.00113	$5.65 \times 10^{-4}$
H	0.01	0.005M	0.00113	$2.825 \times 10^{-4}$
I	0.01	0.005M	0.00113	$1.13 \times 10^{-4}$

(Show calculations for initial concentration data on separate sheet)

In Table 5 below, Equilibrium  $[\text{FeSCN}^{2+}]$  is obtained from the Calibration line using absorbance values of solutions F through I.

Equilibrium  $[\text{Fe}^{3+}] = \text{Initial } [\text{Fe}^{3+}] - \text{Equilibrium } [\text{FeSCN}^{2+}]$

Equilibrium  $[\text{SCN}^-] = \text{Initial } [\text{SCN}^-] - \text{Equilibrium } [\text{FeSCN}^{2+}]$

Table 5. Equilibrium Constant Calculation

Test Tube	Initial $[\text{Fe}^{3+}]$	Initial $[\text{SCN}^-]$	Equilibrium $[\text{FeSCN}^{2+}]$ (from graph)	Equilibrium $[\text{Fe}^{3+}]$	Equilibrium $[\text{SCN}^-]$	$K_c$
F	0.0025	$2.825 \times 10^{-4}$				
G	0.0025	$5.65 \times 10^{-4}$				
H	0.005	$2.825 \times 10^{-4}$				
I	0.005	$1.13 \times 10^{-4}$				

(Show calculations for Equilibrium Constant on separate sheet)

Average Equilibrium Constant \_\_\_\_\_

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