

## Standard Enthalpy of Formation ( $\Delta H^\circ_f$ ) at 298 K

Formula	KJ/mol	Formula	KJ/mol	Formula	KJ/mol
<b>Carbon</b>		<b>Chlorine</b>		<b>Sodium</b>	
C (s, graphite)	0	Cl (g)	121.3	Na (s)	0
C (s, diamond)	1.88	Cl <sub>2</sub> (g)	0	Na (g)	107.5
CH <sub>4</sub> (g)	-74.6	HCl (g)	-92.3	NaCl (s)	-411.2
CH <sub>3</sub> OH (l)	-238.6	<b>Fluorine</b>		Na <sub>2</sub> CO <sub>3</sub> (s)	-1130.7
C <sub>2</sub> H <sub>2</sub> (g)	227.4	F (g)	79.38	NaHCO <sub>3</sub> (s)	-950.8
C <sub>2</sub> H <sub>4</sub> (g)	52.4	F <sub>2</sub> (g)	0	<b>Sulfur</b>	
C <sub>2</sub> H <sub>6</sub> (g)	-84.68	HF (g)	-273.3	S <sub>8</sub> (s, rhombic)	0
C <sub>2</sub> H <sub>5</sub> OH (l)	-277.6	<b>Nitrogen</b>		S <sub>8</sub> (s, monoclinic)	0.3
C <sub>3</sub> H <sub>8</sub> (g)	-103.85	N <sub>2</sub> (g)	0	SO <sub>2</sub> (g)	-296.8
C <sub>3</sub> H <sub>6</sub> O (l, acetone)	-248.4	NH <sub>3</sub> (g)	-45.9	SO <sub>3</sub> (g)	-395.7
<b>Hydrogen</b>		NH <sub>4</sub> (s)	-365.6	H <sub>2</sub> SO <sub>4</sub> (l)	-814.0

# Standard Enthalpy Of Formation For Various Compounds

**David Kirk**



## Standard Enthalpy Of Formation For Various Compounds:

PHYSICAL CHEMISTRY R.C. SARASWAT, S.R. CHIKTE, M.P. SINGH, Unit I Thermodynamics I A Recapitulation of thermodynamic terms System surrounding types of system closed open isolated Thermodynamic variables intensive extensive properties thermodynamic processes isothermal adiabatic isobaric cyclic reversible irreversible State function path functions properties of state functions exact differential cyclic rule integrating factor concept of heat work 3L B Statements of first law of thermodynamics Definition of internal energy enthalpy heat capacity at constant volume at constant pressure Joule Thomson experiment Joule Thomson coefficient Inversion temperature calculations of  $W$   $Q$   $E$   $H$  for expansion of gases for isothermal adiabatic conditions for reversible process carnot s cycle its efficiency thermodynamic scale of temperature 5L C Thermochemistry Heat of reaction standard states relation between heat of reaction at constant volume at constant pressure Hess s law of constant heat of summation its applications bond dissociation energy its calculations from thermochemical data variation of heat of reaction with temperature Kirchoff s equation 4L Unit II Thermodynamics II A Second law of thermodynamics Need for second law of thermodynamics statements of second law of thermodynamics concept of entropy entropy as a state function of  $V$   $T$   $P$   $T$  entropy change in phase change for ideal gas entropy as criteria of spontaneity equilibrium 4L B Free energy functions Helmholtz free energy A Gibb s free energy  $G$  their properties standard free energies effect of temperature on free energy Gibb s Helmholtz equation its applications  $A$   $G$  as criteria for thermodynamic equilibrium 4L C System of variable composition Partial molar quantities chemical potential Van t Hoff s reaction isotherm relation between standard free energy change equilibrium constant thermodynamic derivation of law of mass action effect of temperature on equilibrium constant reaction isochore 4L Unit III Phase Equilibria A Phase rule Statement of phase rule definition of phase component and degree of freedom derivation of phase rule Clapeyron equation its application in deciding slopes of line for two phase equilibria applications of phase rule to two phase equilibria of i water system ii sulphur system iii Pb Ag system 6L B Liquid Liquid mixtures Ideal liquid mixtures Raoult s law of ideal solutions Henry s law non ideal systems azeotropes HCl H<sub>2</sub>O ethanol water system Partial miscible liquids Phenol water system trimethylamine water nicotine water system lower upper consolute temperature effect of impurity Immiscible liquids Steam distillation Nernst distribution law Limitations deviations applications 6L Unit IV Solid State Laws of crystallography i Law of constancy of interfacial angles ii Law of rationality of indices iii Law of symmetry symmetry of elements in crystals Unit cell space lattice orientation of lattice plane Miller indices Bravais lattices crystal systems X ray diffraction by crystal derivation of Braggs equation *Chemistry Class XI - SBPD Publications* Dr. Subhash Chandra Rastogi, ,Er. Meera Goyal, 2022-02-17 Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I

Appendix II Log antilog Table      *Artificial Neural Networks in Chemical Engineering Processes* Ahad Ghaemi,Zohreh Khoshraftar,2025-12-05 *Artificial Neural Networks in Chemical Engineering Processes From Theory to Applications* serves as a comprehensive resource on artificial neural networks within chemical engineering including understanding the fundamental principles learning about relevant algorithms and architectures and exploring practical case studies This book covers theoretical principles relevant algorithms and practical case studies this book covers artificial neural network concepts architectures and algorithms with a focus on applications in chemical engineering processes This book also addressed common challenges by providing practical guidance through successful case studies offering insights on data pre processing model selection training strategies and performance evaluation The book serves as a valuable tool for bridging the gap between neural networks and their practical implementation in chemical engineering This book will be an invaluable resource for chemical Engineers particularly researchers and industry professionals working in Machine Learning and Artificial Intelligence It will also be a very useful guide for Graduate and Postgraduate Students in Chemical Engineering and machine learning *Artificial Neural Networks in Chemical Engineering* will also be a valuable resource for anyone working with artificial neural networks in other industries particularly data scientists and analysts Serves as a comprehensive resource to bridge the gap between theoretical knowledge of neural networks and practical implementation in chemical engineering Provides in depth explanations of neural network concepts architectures and algorithms along with specifics about applications in chemical engineering Outlines various types of artificial neural networks including feed forward networks and their applications in chemical engineering processes and systems Includes practical guidance and case studies that showcase the successful application of neural networks in solving chemical engineering problems Presents insights into essential aspects such as data pre processing techniques model selection training strategies and performance evaluation Provides a roadmap for the effective implementation of neural networks in experimental modeling including code and MATLAB modeling

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